

Chemistry Semester 1 Final Exam Review

1. What is a definition of Chemistry?

Study of matter

2. Matter includes all of the following EXCEPT:

a. air b. light

c. smoke

d. water vapor

3. What is matter?

Something that has mass and takes up space (volume)

4. A physical change occurs when a

a. peach spoils
b. bracelet turns your wrist green

c. copper bowl tarnishes

d. glue gun melts a glue stick

5. The particles in a solid are

a. packed closely together
b. very far apart

c. constantly in motion

d. able to slide past one another

Of the following groups, 1, 2, 14, 15, 17, 18:

6. Group(s) **1, 2** contains only metals

7. Group(s) **18** is called a Noble gases

8. Group(s) **17** has 7 electrons on the valence shell

9. Group(s) **15** needs to gain 3 electrons to fill a stable octet

10. Group(s) **2** has a 2+ oxidation state

11. Group(s) **2** has an electron dot notation of ***X***

12. Using the periodic table, an element similar to carbon would be **14**

13. Based on their location on the periodic table, you could infer that **18/noble gases** are very unreactive

14. The horizontal row on the periodic table is called a(n) **period**

15. The vertical column of blocks on the periodic table is called a(n) **group/family**

16. The elements on the zigzag line in the periodic table are **metalloids**

17. Which is NOT a property of metal?

a. malleability
b. ability to lose electrons

c. unreactivity

d. ability to conduct heat

18. A volume of 1 cc (cm³) is the same as **1** ml

19. To determine density, the quantities that must be measured are **volume** and **mass**

20. The relationship between the mass *m* of a material, its volume *V*, and its density *D* is

a. $V = mD$

c. $DV = m$

d. $D + V = m$

21. The density of aluminum is 2.70 g/cm³. The volume of a solid piece of Aluminum is 1.50 cm³. Find the mass **4.05 grams**

22. The density of pure diamond is 3.5 g/cm³. The mass of a diamond is 0.25 g. Find the volume **0.71 cm³**

22. What statement about density is true?

a. two samples of a pure substance may have different densities if they are different sizes
b. the density of a sample depends on its location on Earth
c. a cylinder is always used to measure the volume
d. density is a physical property that remains constant for a pure substance

23. The number of grams equal to 0.5 kg is **500 g**

24. In a graph the **independent** variable is placed on the x axis and the **dependant** variable is placed on the y axis.

25. The variable that you change in the lab is the **independent** variable. The variable that responds to the change is called the **dependant** variable. All other things do not change and are called **constants**

26. At first Dalton thought the atom was **solid, spherical, indivisible**, then Rutherford discovered the **nucleus** of the atom, then the subatomic particles **protons, neutrons, electrons** were discovered

27. The positively charged particle is the **proton** found in the **nucleus** of the atom

28. The subatomic particle that has about the same mass as the proton, but with no electrical charge, is called a **neutron** and is found in the **nucleus** of the atom

29. The **electron** is negatively charged and is found in the **electron cloud/outer shell** of the atom

30. **Bloch** is the person credited with placing electrons on levels and **Schrodinger** established the electron cloud theory.

31. The forces that hold the particles together in the nucleus are **coulombic/electrostatic** forces

32. Isotopes are atoms of the same element that **have different number of neutrons**

33. All isotopes of hydrogen contain **1 proton & 1 electron**

34. What is the atomic number for Nitrogen? 7 To what subatomic particle does this correspond? **# of protons**

35. Zinc-66 has **30** protons and **36** neutrons

36. What is the electron configuration for sodium? **[Ne]3s¹** Chlorine? **[Ne]3s²3p⁵** Neon? **[He] 2s²2p⁴**

37. What is the dot notation for magnesium? Bromine? Argon?

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