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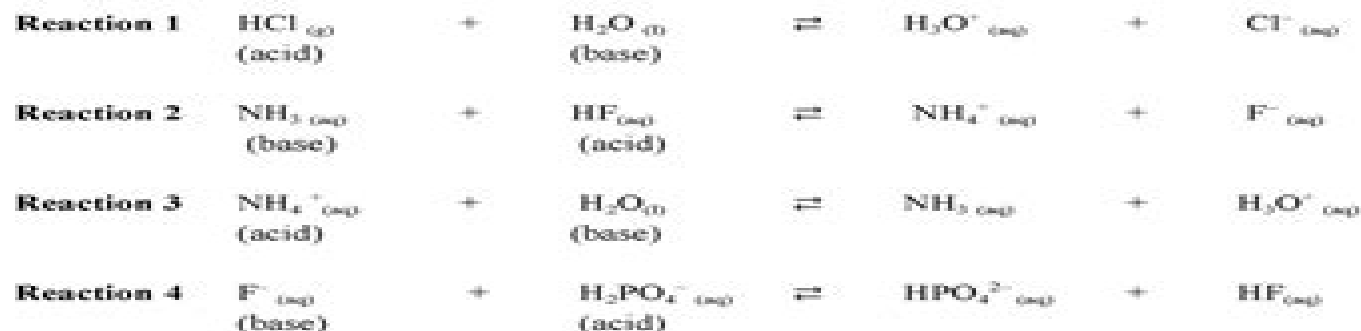
## Acids and Bases

How do acids and bases behave in water?

### Why?

Acids and bases play an important role in our lives. Numerous biological processes, industrial applications, and even environmental problems are a function of the acidity or basicity (alkalinity) of aqueous solutions. It is therefore important to understand what makes a substance behave as an acid or a base when dissolved in water. In this activity, we will explore the physical and chemical properties of acids and bases.

### Model 2 - Bronsted-Lowry Acids and Bases



1. Identify the Bronsted-Lowry acids in Model 2.

a. Atoms of which element are present in all of the Bronsted-Lowry acids in Model 2?

Hydrogen

b. How can you tell from Reaction 1 that HCl loses an  $\text{H}^+$  ion rather than a hydrogen atom when the reaction occurs. **Hint:** Look at the products.

The Cl has a negative charge. The positive from the hydrogen and the negative from the chloride cancel each other out when in a compound.

2. For each acid-base reaction in Model 2, describe the role of the Bronsted-Lowry acid in the  $\text{H}^+$  ion (proton) transfer that occurs.

The acids lose a hydrogen ion

# Acids Base Answer Key Pogil

**Adam Hulanicki**



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*Analytical Chemistry* Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more

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Acid-bases in Analytical Chemistry Izaak Maurits Kolthoff, 1966      **The ABC of Acid-base Chemistry ...** Horace Willard Davenport, 1974      **Reactions of Acids and Bases in Analytical Chemistry** Adam Hulanicki, 1987      A New View of Current Acid-base Theories H. L. Finston, Allen C. Rychtmann, 1982      *Acid-Base Diagrams* Heike Kahlert, Fritz

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**Proton Chemistry** Robert A. Richardson,1976  
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*Acid Bases in Analytical Chemistry* Izaak Maurits Kolthoff,S. Bruckenstein,1964-01-01

*Acids and Bases* Brian G. Cox,2013-01-31 Acids and bases are ubiquitous in chemistry Our understanding of them however is dominated by their behaviour in water Transfer to non aqueous solvents leads to profound changes in acid base strengths and to the rates and equilibria of many processes for example synthetic reactions involving acids bases and nucleophiles isolation of pharmaceutical actives through salt formation formation of zwitter ions in amino acids and chromatographic separation of substrates This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non aqueous solvents The behaviour is related where possible to that in water but correlations and contrasts between solvents are also presented Fundamental background material is provided in the initial chapters quantitative aspects of acid base equilibria including definitions and relationships between solution pH and species distribution the influence of molecular structure on acid strengths and acidity in aqueous solution Solvent properties are reviewed along with the magnitude of the interaction energies of solvent molecules with especially ions the ability of solvents to participate in hydrogen bonding and to accept or donate electron pairs is seen to be crucial Experimental methods for determining dissociation constants are described in detail In the remaining chapters dissociation constants of a wide range of acids in three distinct classes of solvents are discussed protic solvents such as alcohols which are strong hydrogen bond donors basic polar aprotic solvents such as dimethylformamide and low basicity and low polarity solvents such as acetonitrile and tetrahydrofuran Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents and there is a strong differentiation between the response of neutral and charged acids to solvent change Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described

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