

## Chemistry 2<sup>nd</sup> Semester Final Exam Review

### Chemical Bonds

1. Give a physical description of how the atoms and molecules are arranged in solids, liquids, and gases.  
*A: In a liquid, the forces between the molecules are weak, so the molecules are free to move around. In a solid the forces between the molecules (intermolecular forces) are stronger so they are held in place.*
2. Use the information from the previous question to describe the shape and volume for each phase (state) of matter.  
*A: Solids have a repeating geometric pattern (lattice structure) in which the particles vibrate in place. They have a definite shape and a definite volume.  
Particles within a liquid slide past each and are close, but are not held in place. They have a definite volume, but no definite shape.  
Gases particles move chaotically bouncing off one another. They have neither a definite shape nor a definite volume.*
3. List the three phases of matter in order of increasing intermolecular attractions.  
*A: Gas, liquid, solid.*
4. Why do the atoms and molecules in liquids move in a random pattern relative to one another instead of being in a solid form?  
*A: Their kinetic energy (temperature) is high enough to overcome the attractive forces that might otherwise hold them in a repeating pattern. The attractive forces aren't strong enough to hold them together at that temperature.*

### Conservation of Matter and Stoichiometry

5. The atomic mass unit and the mole are defined using what isotope? *A: Carbon-12*
6. Balance the following equations:
  - a.  $\text{Al}_2\text{O}_3 + \text{Cl}_2 + \text{C} \rightarrow \text{AlCl}_3 + \text{CO}$   
*A:  $\text{Al}_2\text{O}_3 + 3\text{Cl}_2 + 3\text{C} \rightarrow 2\text{AlCl}_3 + 3\text{CO}$*
  - b.  $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$   
*A:  $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$*
  - c.  $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$   
*A:  $4\text{P} + 5\text{O}_2 \rightarrow 2\text{P}_2\text{O}_5$*
7. What isotope is used to define one mole? *A: Carbon-12*
8. How many particles (atoms or molecules) are in one mole? *A:  $6.02 \times 10^{23}$*
9. Find the molar mass for the following chemical formulas:
  - a.  $\text{ZnF}_2$  *A: 103.377 g/mol*
  - b.  $\text{Al}_2(\text{SO}_4)_3$  *A: 342.151 g/mol*
  - c.  $\text{NH}_4\text{IO}_3$  *A: 192.941 g/mol*
10. How many moles are in 4.37 kg of NaOH? *A: 109.3 mol*
11. How many atoms are in 2075 g of He? *A:  $3.12 \times 10^{26}$  atoms*
12. What volume will 92.31 g of  $\text{CO}_2$  occupy at STP? *A: 46.985 L*
13. How many grams of sodium hydroxide are needed to react with 6.23 g of barium bromide?  
 $\text{NaOH} + \text{BaBr}_2 \rightarrow \text{NaBr} + \text{Ba(OH)}_2$   
*A: 1.68 g NaOH*
14. In the reaction,  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ , if 100.0 g of magnesium reacts with 50.0 g of oxygen, what mass of product is produced?  
*A: 126 g MgO*

### Gases and Their Properties

15. What is the kinetic molecular theory (KMT)?  
*A: The Kinetic Molecular Theory is used to explain the behavior of gases and is based upon the following postulates:*
  - Gases are composed of a many particles that behave like hard spherical objects in a state of constant, random motion.
  - These particles move in a straight line until they collide with another particle or the walls of the container.
  - These particles are much smaller than the distance between particles, therefore the volume of a gas is mostly empty space and the volume of the gas molecule themselves is negligible.
  - There is no force of attraction between gas particles or between the particles and the walls of the container.
  - Collisions between gas particles or collisions with the walls of the container are elastic. That is, none of the energy of the gas particle is lost in a collision.

# Nd Semester Final Review Questions

**Guillaume Favre**



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**Proceedings of the Board of Public Instruction of the City of Albany**, 1885

*How to Study Better and Get Higher Marks* Eugene Ehrlich, 1976

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*PSEA Education Bulletin*, 1937

*Course of Study for the Public Schools of Eau Claire, Wisconsin, for the Grades Eau Claire (Wis.). Board of Education, 1919*

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