

Honors Chemistry
Extra Stoichiometry Problems

1. Silver nitrate reacts with barium chloride to form silver chloride and barium nitrate.

a. Write and balance the chemical equation.



b. If 39.02 grams of barium chloride are reacted in an excess of silver nitrate, how many representative particles (and what type) of silver chloride are produced?

$$\frac{39.02 \text{ g BaCl}_2}{1} \times \frac{1 \text{ mole BaCl}_2}{208.233 \text{ g BaCl}_2} \times \frac{2 \text{ moles AgCl}}{1 \text{ mole BaCl}_2} \times \frac{6.02 \times 10^{23} \text{ f. u. AgCl}}{1 \text{ mole AgCl}} = 2.256 \times 10^{23} \text{ f. u. AgCl}$$

c. If 410.8 grams of barium nitrate are produced how many grams of silver nitrate were reacted?

$$\frac{410.8 \text{ g Ba}(\text{NO}_3)_2}{1} \times \frac{1 \text{ mole Ba}(\text{NO}_3)_2}{261.335 \text{ g Ba}(\text{NO}_3)_2} \times \frac{2 \text{ moles AgNO}_3}{1 \text{ mole Ba}(\text{NO}_3)_2} \times \frac{169.872 \text{ g AgNO}_3}{1 \text{ mole AgNO}_3} = 534.1 \text{ g AgNO}_3$$

2. Nitrogen gas is reacted with hydrogen gas to form nitrogen trihydride.

a. Write and balance the chemical equation.



b. How many liters of nitrogen trihydride are produced at STP if 80.28 grams of hydrogen gas are reacted in an excess of nitrogen?

$$\frac{80.28 \text{ g H}_2}{1} \times \frac{1 \text{ mole H}_2}{2.016 \text{ g H}_2} \times \frac{2 \text{ moles NH}_3}{3 \text{ moles H}_2} \times \frac{22.4 \text{ L NH}_3}{1 \text{ mole NH}_3} = 594.7 \text{ L NH}_3$$

c. How many grams of hydrogen are needed to fully react 621.9 grams of nitrogen gas?

$$\frac{621.9 \text{ g N}_2}{1} \times \frac{1 \text{ mole N}_2}{28.014 \text{ g N}_2} \times \frac{3 \text{ moles H}_2}{1 \text{ mole N}_2} \times \frac{2.016 \text{ g H}_2}{1 \text{ mole H}_2} = 134.3 \text{ g H}_2$$

d. How many liters of nitrogen gas are needed to completely react 90.38 L of hydrogen gas at STP?

$$\frac{90.38 \text{ L H}_2}{1} \times \frac{1 \text{ mole H}_2}{22.4 \text{ L H}_2} \times \frac{1 \text{ mole N}_2}{3 \text{ mole H}_2} \times \frac{22.4 \text{ L N}_2}{1 \text{ mole N}_2} = 30.09 \text{ L N}_2$$

OR

$$\frac{90.38 \text{ L H}_2}{1} \times \frac{1 \text{ L N}_2}{3 \text{ L H}_2} = 30.09 \text{ L N}_2$$

12 Stoichiometry Practice Problems Answer Key

James E. Brady, Gerard E. Humiston

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Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book. It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features. Each section focuses on the development and assessment of one or two specific objectives. Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives. These features are followed by a set of independent practice problems. The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics. Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance.

Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change. *Introduction to Chemistry, Laboratory Manual* T. R. Dickson, 1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives. Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities. These themes involve and engage the students in their own learning processes they are challenged to be active. The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics. The reorganization presents many core concepts in the first five chapters preparing students for later chapters. In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students.

Stoichiometry Unit Project Luann Marie Decker, 1998 **Chemistry** John McMurry, Robert C. Fay, 2001 A colorful pedagogically enhanced standard textbook for the introductory course. It begins with atomic structure proceeds next to bonding and molecules then to bulk physical properties of substances and ends with a study of chemical properties. Each chapter concludes with a brief description of an interesting application or extension of the chapter subject a summary a list of key words and a large number of problems. Many student oriented supplements are available. Annotation copyright by Book News Inc Portland OR.

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Addison-Wesley Chemistry, 1990 *Instructors Resource Manual* John McMurry, Robert A. Pribush, Robert Fay, 2000-12-27 **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Resources in Education*, 1980-06 Chemistry Success in 20 Minutes a Day Michael B. McGinnis, 2005 Offers a diagnostic test and twenty lessons covering vital chemistry skills.

The Software Encyclopedia 2000 Bowker Editorial Staff, 2000-05

Chemistry Charles H. Corwin, 1994 The book focuses on the concepts of chemistry and the applications that maintain and generate motivation for the subject of chemistry. General Chemistry Kenneth W. Whitten, Kenneth D. Gailey, Raymond

E. Davis,1988 **Introductory Chemistry** Steve Russo,Mike Silver,2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support onThe Chemistry Place tm Special Edition and onThe Chemistry of Life CD ROM For college instructors and students *General Chemistry*
Ralph Sherman Becker,Wayne Ernest Wentworth,1980 **General Chemistry** James E. Brady,Gerard E. Humiston,1980

General Chemistry with Qualitative Analysis Ralph H. Petrucci,Robert K. Wismer,1987 **Treatise on Analytical Chemistry: Theory and practice.** v Izaak Maurits Kolthoff,Philip Juliber Elving,1959 **Children's Books in Print** R R Bowker Publishing,Bowker,1999-12

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Table of Contents 12 Stoichiometry Practice Problems Answer Key

1. Understanding the eBook 12 Stoichiometry Practice Problems Answer Key
 - The Rise of Digital Reading 12 Stoichiometry Practice Problems Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 12 Stoichiometry Practice Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 12 Stoichiometry Practice Problems Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 12 Stoichiometry Practice Problems Answer Key
 - Personalized Recommendations
 - 12 Stoichiometry Practice Problems Answer Key User Reviews and Ratings
 - 12 Stoichiometry Practice Problems Answer Key and Bestseller Lists
5. Accessing 12 Stoichiometry Practice Problems Answer Key Free and Paid eBooks

- 12 Stoichiometry Practice Problems Answer Key Public Domain eBooks
- 12 Stoichiometry Practice Problems Answer Key eBook Subscription Services
- 12 Stoichiometry Practice Problems Answer Key Budget-Friendly Options

6. Navigating 12 Stoichiometry Practice Problems Answer Key eBook Formats

- ePUB, PDF, MOBI, and More
- 12 Stoichiometry Practice Problems Answer Key Compatibility with Devices
- 12 Stoichiometry Practice Problems Answer Key Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 12 Stoichiometry Practice Problems Answer Key
- Highlighting and Note-Taking 12 Stoichiometry Practice Problems Answer Key
- Interactive Elements 12 Stoichiometry Practice Problems Answer Key

8. Staying Engaged with 12 Stoichiometry Practice Problems Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 12 Stoichiometry Practice Problems Answer Key

9. Balancing eBooks and Physical Books 12 Stoichiometry Practice Problems Answer Key

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 12 Stoichiometry Practice Problems Answer Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 12 Stoichiometry Practice Problems Answer Key

- Setting Reading Goals 12 Stoichiometry Practice Problems Answer Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 12 Stoichiometry Practice Problems Answer Key

- Fact-Checking eBook Content of 12 Stoichiometry Practice Problems Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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