

# Assigning Oxidation Numbers

Rule	Example
1. The oxidation number of any uncombined element is 0.	The oxidation number of $\text{Na}(s)$ is 0.
2. The oxidation number of a monatomic ion equals the charge on the ion.	The oxidation number of $\text{Cl}^-$ is -1.
3. The more-electronegative element in a binary compound is assigned the number equal to the charge it would have if it were an ion.	The oxidation number of O in $\text{NO}$ is -2.
4. The oxidation number of fluorine in a compound is always -1.	The oxidation number of F in $\text{LiF}$ is -1.
5. Oxygen has an oxidation number of -2 unless it is combined with F, in which it is +1 or +2, or it is in a peroxide, in which it is -1.	The oxidation number of O in $\text{NO}_2$ is -2.
6. Hydrogen's oxidation state in most of its compounds is +1 unless it is combined with a metal, in which case it is -1.	The oxidation number of H in $\text{LiH}$ is -1.
7. In compounds, the elements of Groups 1 and 2 as well as aluminum have oxidation numbers of +1, +2, and +3, respectively.	The oxidation number of Ca in $\text{CaCO}_3$ is +2.
8. The sum of the oxidation numbers of all atoms in a neutral compound is 0.	The oxidation number of C in $\text{CaCO}_3$ is +4.
9. The sum of the oxidation numbers of all atoms in a polyatomic ion equals the charge of the ion.	The oxidation number of P in $\text{H}_2\text{PO}_4^-$ is +5.

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**Ibrahim Sikder**

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