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web limiting reactant theoretical and percentage yields key $4 \text{ KO}_2 \text{ s } 2 \text{ H}_2 \text{ O l } 4 \text{ KOH aq } 3 \text{ O}_2 \text{ g}$ if a reaction vessel contains 0 mol KO_2 and 0 mol $\text{H}_2 \text{ O}$ what is the limiting reactant

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 $\text{Cu} + 2 \text{ AgNO}_3 \rightarrow 2 \text{ Ag} + \text{Cu}(\text{NO}_3)_2$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 302.3 \text{ g}$ c determine the percent yield $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7 $\text{Na}_2\text{CO}_3 + 4 \text{HCl} \rightarrow 2 \text{NaCl} + \text{CO}_2 + \text{H}_2\text{O}$ is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H_2 are left over 13

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