

Experiment 6: Activity Series: Single Replacement Reactions

Required reading: Ebbing, 11th Edition Chapter 4.

- Molecular and Ionic equations
- Displacement reactions
- Balancing redox equations

Learning Goals:

- To observe reactions of metals with acids and salt solutions using activity series of metals.
- To be able to write and balance single replacement reactions using activity series of metals.
- To be able to write ionic and net ionic equations for the above reactions.

Background information and theory:

One of the most important characteristics of a metal is its activity (reactivity). The activity of a metal is its ability to react with nonmetals. In such reaction, a metal typically loses electrons to the nonmetal to form cations and the product is an ionic compound. Different metals have different activity. Some metals are very active such as lithium, sodium, potassium, and cesium; some are slowly active such as tin, lead, and iron; and some are not active at all (inactive) such as copper, silver, and mercury. Inactive metals are those which resist oxidation. These are often called the noble metals.

An **activity series** is the ranking of metals according to their reactivity. In an activity series metals are arranged in order of decreasing ability to lose electrons. Generally, the more reactive metals are placed at the left of the list.

Li > K > Ba > Ca > Na > Mg > Al > Mn > Zn > Cr > Fe > Co > Ni > Sn > Pb > H > Cu > Ag > Hg > Pt > Au
More active

→

 less active

Any metal before hydrogen will displace hydrogen gas, $H_{2(g)}$, from an acid, and any metal after hydrogen will not displace hydrogen from an acid. For example:

Active metals (left to H) : $Zn_{(s)} + 2HCl_{(aq)} \rightarrow ZnCl_{2(aq)} + H_{2(g)}$
Inactive metals (right to H) : $Cu_{(s)} + HCl_{(aq)} \rightarrow \text{no reaction (NR)}$

In performing this experiment, you will observe and compare the reaction of some common metals with acids and salt solutions and establish an order of reactivity, from most active to least active.

PART I: Reactivity of metals with HCl

Very active metals will react with strong acids such as hydrochloric acid. An example considers magnesium:

Molecular equation: $Mg_{(s)} + 2HCl_{(aq)} \rightarrow MgCl_{2(aq)} + H_{2(g)}$

Activity Series Lab Experiment 15 Answers

Huangqi Zhang



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Table of Contents Activity Series Lab Experiment 15 Answers

1. Understanding the eBook Activity Series Lab Experiment 15 Answers
 - The Rise of Digital Reading Activity Series Lab Experiment 15 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Activity Series Lab Experiment 15 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Activity Series Lab Experiment 15 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Activity Series Lab Experiment 15 Answers
 - Personalized Recommendations
 - Activity Series Lab Experiment 15 Answers User Reviews and Ratings
 - Activity Series Lab Experiment 15 Answers and Bestseller Lists
5. Accessing Activity Series Lab Experiment 15 Answers Free and Paid eBooks

- Activity Series Lab Experiment 15 Answers Public Domain eBooks
- Activity Series Lab Experiment 15 Answers eBook Subscription Services
- Activity Series Lab Experiment 15 Answers Budget-Friendly Options
- 6. Navigating Activity Series Lab Experiment 15 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Activity Series Lab Experiment 15 Answers Compatibility with Devices
 - Activity Series Lab Experiment 15 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Activity Series Lab Experiment 15 Answers
 - Highlighting and Note-Taking Activity Series Lab Experiment 15 Answers
 - Interactive Elements Activity Series Lab Experiment 15 Answers
- 8. Staying Engaged with Activity Series Lab Experiment 15 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Activity Series Lab Experiment 15 Answers
- 9. Balancing eBooks and Physical Books Activity Series Lab Experiment 15 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Activity Series Lab Experiment 15 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Activity Series Lab Experiment 15 Answers
 - Setting Reading Goals Activity Series Lab Experiment 15 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Activity Series Lab Experiment 15 Answers
 - Fact-Checking eBook Content of Activity Series Lab Experiment 15 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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