

**POGIL Activity: Lewis Structures and Drawing Bonds****Why? READ THIS!**

We will learn there are polar and nonpolar covalent compounds. In **polar covalent** compounds there is an **unequal** sharing of electrons between atoms, whereas in **nonpolar covalent** compounds there is an **equal/almost equal** sharing of electrons between atoms. In addition to determining polarity based upon electronegativities, we can determine the polarity based upon structural drawings.

The polarity of a compound can tell us a lot about the properties of a certain compound, so learning how to determine the polarity will be helpful. However, we must first learn how to draw the compounds in order to determine bond polarity. In addition, we can take the structural formula of the compound and determine the 3-D shape of the compound.

**Model 1: Lewis Structures of Elements**

You learned about Lewis structures in the previous chapter. Lewis structures are used to represent the **valence electrons** around the element. You must be able to draw the Lewis structure of any main group element before you can draw the Lewis structure of a compound.

Again, here are the steps for drawing Lewis structures (Electron Dot Diagram) for elements:

1. Write the symbol for the element
2. Determine the number of valence electrons the element has
3. Draw the valence electrons as dots surrounding the element
  - a. We designate 4 sides to the element symbol: top, right, bottom, left
  - b. Before you can put two dots on the same side, each side must have one dot (2 max)

**Questions:**

1. Draw the Lewis structures (Electron Dot Diagram) for the following elements:

a. Ca	b. P	c. H	d. C	e. Cl
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**Model 2: Lewis Structures and Structural Formulas of Single Bonds**

Now that you remember how to draw Lewis structures for individual elements, let's draw the Lewis structures for compounds. The Lewis structures of compounds use straight lines to represent the two shared electrons in a bond. The structural formulas can then be used to draw the 3-D structures of covalent compounds.

In covalent compounds with **single bonds**, there is only one shared pair of electrons between the central atom and the other atoms.

Example: Water (H<sub>2</sub>O)

Molecular Formula	Electron Dot Diagram	Lewis Structure
H <sub>2</sub> O	$\begin{array}{c} \text{H} \cdot \quad \text{H} \cdot \quad \cdot \ddot{\text{O}} : \\ \text{Total Valence Electrons} = 8 \end{array}$	$\begin{array}{c} \text{H} : \ddot{\text{O}} : \\   \\ \text{H} \end{array} \quad \begin{array}{c} \text{H} - \ddot{\text{O}} : \\   \\ \text{H} \end{array}$

**NOTE: Hydrogen is full at 2 valence electrons!**

# Activity In Drawing Lewis Structures Pogil

**Suzanne M. Ruder, The POGIL Project**



### **Activity In Drawing Lewis Structures Pogil:**

**Organic Chemistry** Suzanne M. Ruder, The POGIL Project, 2015-12-29 ORGANIC CHEMISTRY *Cooperative Learning in Higher Education* Barbara Millis, 2023-07-03 Research has identified cooperative learning as one of the ten High Impact Practices that improve student learning If you've been interested in cooperative learning but wondered how it would work in your discipline this book provides the necessary theory and a wide range of concrete examples Experienced users of cooperative learning demonstrate how they use it in settings as varied as a developmental mathematics course at a community college and graduate courses in history and the sciences and how it works in small and large classes as well as in hybrid and online environments The authors describe the application of cooperative learning in biology economics educational psychology financial accounting general chemistry and literature at remedial introductory and graduate levels The chapters showcase cooperative learning in action at the same time introducing the reader to major principles such as individual accountability positive interdependence heterogeneous teams group processing and social or leadership skills The authors build upon and cross reference each others chapters describing particular methods and activities in detail They explain how and why they may differ about specific practices while exemplifying reflective approaches to teaching that never fail to address important assessment issues

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