

**POGIL Activity: Lewis Structures and Drawing Bonds****Why? READ THIS!**

We will learn there are polar and nonpolar covalent compounds. In **polar covalent** compounds there is an **unequal** sharing of electrons between atoms, whereas in **nonpolar covalent** compounds there is an **equal/almost equal** sharing of electrons between atoms. In addition to determining polarity based upon electronegativities, we can determine the polarity based upon structural drawings.

The polarity of a compound can tell us a lot about the properties of a certain compound, so learning how to determine the polarity will be helpful. However, we must first learn how to draw the compounds in order to determine bond polarity. In addition, we can take the structural formula of the compound and determine the 3-D shape of the compound.

**Model 1: Lewis Structures of Elements**

You learned about Lewis structures in the previous chapter. Lewis structures are used to represent the **valence electrons** around the element. You must be able to draw the Lewis structure of any main group element before you can draw the Lewis structure of a compound.

Again, here are the steps for drawing Lewis structures (Electron Dot Diagram) for elements:

1. Write the symbol for the element
2. Determine the number of valence electrons the element has
3. Draw the valence electrons as dots surrounding the element
  - a. We designate 4 sides to the element symbol: top, right, bottom, left
  - b. Before you can put two dots on the same side, each side must have one dot (2 max)

**Questions:**

1. Draw the Lewis structures (Electron Dot Diagram) for the following elements:

a. Ca	b. P	c. H	d. C	e. Cl
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**Model 2: Lewis Structures and Structural Formulas of Single Bonds**

Now that you remember how to draw Lewis structures for individual elements, let's draw the Lewis structures for compounds. The Lewis structures of compounds use straight lines to represent the two shared electrons in a bond. The structural formulas can then be used to draw the 3-D structures of covalent compounds.

In covalent compounds with **single bonds**, there is only one shared pair of electrons between the central atom and the other atoms.

Example: Water ( $\text{H}_2\text{O}$ )

Molecular Formula	Electron Dot Diagram	Lewis Structure
$\text{H}_2\text{O}$	$\text{H}^+$ $\text{H}^+$ $\cdot\ddot{\text{O}}:$ Total Valence Electrons = 8	$\text{H}:\ddot{\text{O}}:$ $\text{H}-\ddot{\text{O}}:$ $\text{H}$ $\text{H}$

**NOTE: Hydrogen is full at 2 valence electrons!**

# Activity In Drawing Lewis Structures Pogil

**Suzanne M. Ruder, The POGIL Project**

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**Organic Chemistry** Suzanne M. Ruder, The POGIL Project, 2015-12-29 ORGANIC CHEMISTRY      *Cooperative Learning in Higher Education* Barbara Millis, 2023-07-03 Research has identified cooperative learning as one of the ten High Impact Practices that improve student learning. If you've been interested in cooperative learning but wondered how it would work in your discipline this book provides the necessary theory and a wide range of concrete examples. Experienced users of cooperative learning demonstrate how they use it in settings as varied as a developmental mathematics course at a community college and graduate courses in history and the sciences and how it works in small and large classes as well as in hybrid and online environments. The authors describe the application of cooperative learning in biology, economics, educational psychology, financial accounting, general chemistry and literature at remedial, introductory and graduate levels. The chapters showcase cooperative learning in action at the same time introducing the reader to major principles such as individual accountability, positive interdependence, heterogeneous teams, group processing and social or leadership skills. The authors build upon and cross reference each other's chapters, describing particular methods and activities in detail. They explain how and why they may differ about specific practices while exemplifying reflective approaches to teaching that never fail to address important assessment issues.

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## **Table of Contents Activity In Drawing Lewis Structures Pogil**

1. Understanding the eBook Activity In Drawing Lewis Structures Pogil
  - The Rise of Digital Reading Activity In Drawing Lewis Structures Pogil
  - Advantages of eBooks Over Traditional Books
2. Identifying Activity In Drawing Lewis Structures Pogil
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Activity In Drawing Lewis Structures Pogil
  - User-Friendly Interface
4. Exploring eBook Recommendations from Activity In Drawing Lewis Structures Pogil
  - Personalized Recommendations
  - Activity In Drawing Lewis Structures Pogil User Reviews and Ratings
  - Activity In Drawing Lewis Structures Pogil and Bestseller Lists
5. Accessing Activity In Drawing Lewis Structures Pogil Free and Paid eBooks
  - Activity In Drawing Lewis Structures Pogil Public Domain eBooks
  - Activity In Drawing Lewis Structures Pogil eBook Subscription Services
  - Activity In Drawing Lewis Structures Pogil Budget-Friendly Options

### 6. Navigating Activity In Drawing Lewis Structures Pogil eBook Formats

- ePUB, PDF, MOBI, and More
- Activity In Drawing Lewis Structures Pogil Compatibility with Devices
- Activity In Drawing Lewis Structures Pogil Enhanced eBook Features

### 7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Activity In Drawing Lewis Structures Pogil
- Highlighting and Note-Taking Activity In Drawing Lewis Structures Pogil
- Interactive Elements Activity In Drawing Lewis Structures Pogil

### 8. Staying Engaged with Activity In Drawing Lewis Structures Pogil

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Activity In Drawing Lewis Structures Pogil

### 9. Balancing eBooks and Physical Books Activity In Drawing Lewis Structures Pogil

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Activity In Drawing Lewis Structures Pogil

### 10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

### 11. Cultivating a Reading Routine Activity In Drawing Lewis Structures Pogil

- Setting Reading Goals Activity In Drawing Lewis Structures Pogil
- Carving Out Dedicated Reading Time

### 12. Sourcing Reliable Information of Activity In Drawing Lewis Structures Pogil

- Fact-Checking eBook Content of Activity In Drawing Lewis Structures Pogil
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements

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- Interactive and Gamified eBooks

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