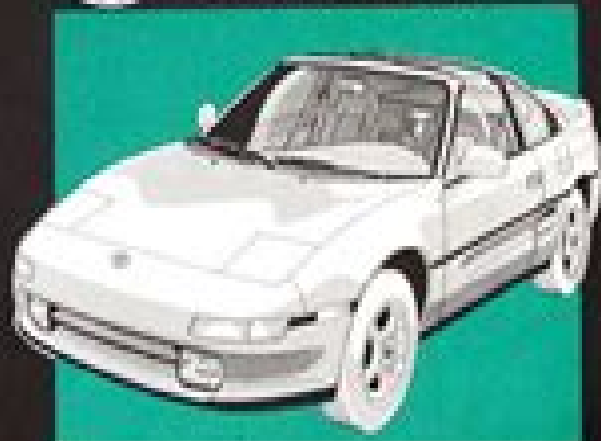


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constant for this reaction at the given temperature of 700 c
 $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
 i write the initial concentrations of reactants and products

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$$\Delta G = \Delta G^\circ + RT \ln K_p$$

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$$\text{H}_2(\text{g}) + 2\text{HI}(\text{g}) \rightleftharpoons 2\text{H}(\text{g}) + \text{I}_2(\text{g})$$

 at 425°C an equilibrium mixture has the following concentrations
 $[\text{H}_2] = 0.1 \times 10^{-2} \text{ mol l}^{-1}$ $[\text{HI}] = 2.125 \times 10^{-3} \text{ mol l}^{-1}$ $[\text{I}_2] = 1.49 \times 10^{-3} \text{ mol l}^{-1}$

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