

# Acid-Base Equilibrium Practice

1. What is the molarity of a solution of nitric acid if 0.216g of barium hydroxide is required to neutralize 20.0ml of nitric acid?
2. What is the volume of 1.222M sodium hydroxide required to react with:
  - a. 32.5ml of 0.569M sulfuric acid
  - b. 5.00g of oxalic acid,  $\text{H}_2\text{C}_2\text{O}_4$ ?
  - c. 15.0g of concentrated acetic acid?
3. What is the volume of 0.885M hydrochloric acid required to react with:
  - a. 25.00ml of 0.288M aqueous ammonia?
  - b. 10.0g of sodium hydroxide?
  - c. 25.0ml of a solution ( $d=0.928\text{g/cm}^3$ ) containing 10.0% by mass of methylamine,  $\text{CH}_3\text{NH}_2$ ?
4. A capsule of vitamin C, a weak acid, is analyzed by titrating it with a 0.425M sodium hydroxide solution. It is found that 6.20ml of bas is required to react with a capsule weighing 0.628g. What is the percentage of vitamin C,  $\text{C}_6\text{H}_8\text{O}_6$ , in the capsule?
5. An artificial beverage contains 12.0g of tartaric acid,  $\text{H}_2\text{C}_4\text{H}_4\text{O}_6$ , to achieve tartness. It is titrated with a basic solution that has a density of 1.045g/cm<sup>3</sup> and contains 5% by mass, KOH. What volume of the basic solution is required?
6. Lactic acid,  $\text{C}_3\text{H}_6\text{O}_3$ , is the acid in sour milk. A 0.100g sample of pure lactic acid requires 12.95ml of 0.0857M sodium hydroxide for complete reaction. How many moles of hydroxide ions are required to neutralize one mole of lactic acid?
7. An aqueous solution contains dissolved  $\text{C}_6\text{H}_5\text{NH}_2\text{Cl}$  and  $\text{C}_6\text{H}_5\text{NH}_3^+$ . The concentration of  $\text{C}_6\text{H}_5\text{NH}_3^+$  is 0.50M and a pH of 4.20.  $K_b = 3.8 \times 10^{-10}$ 
  - a. Calculate the concentration of  $\text{C}_6\text{H}_5\text{NH}_2^+$  in this buffer solution.
  - b. Calculate the pH after 4.0grams of HCl is added to 1.0L of solution. Assume no volume change.
8. A buffer solution is made up of 355ml each of 0.200M  $\text{NaHCO}_3$  and 0.134M  $\text{Na}_2\text{CO}_3$ , calculate:  $\text{ka} = 5.6 \times 10^{-11}$ 
  - a. The pH of the buffer
  - b. The pH of the buffer after the addition of 0.0300mol HCl to 0.710L of buffer solution
  - c. The pH of the buffer after the addition of 0.0300mol KOH to 0.710L of buffer solution
9. Calculate the pH of a solution prepared by mixing 100.0ml of 1.20M ethanolamine,  $\text{C}_2\text{H}_5\text{ONH}_2$ , with 50.0ml of 1.0M HCl.  $\text{Ka}$  for  $\text{C}_2\text{H}_5\text{ONH}_3^+ = 6.3 \times 10^{-10}$ .
10. Calculate the pH of the following solution:
  - a. 2.0M  $\text{NaC}_2\text{H}_5\text{O}_2$  / 2.0M  $\text{HC}_2\text{H}_5\text{O}_2$
  - b. The pH of the buffer solution after 50.0ml of 0.15M LiOH has been added to 100.0ml of the buffer solution.

# Acid Base Equilibrium Practice Problems Multiple Choice

**Peter J. Mikulecky, Michelle Rose  
Gilman, Kate Brutlag**

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