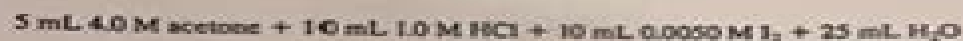


Experiment 20

Advance Study Assignment: The Iodination of Acetone

- In a reaction involving the iodination of acetone, the following volumes were used to make up the reaction mixture:



- How many moles of acetone were in the reaction mixture? Recall that, for a component A, moles $A = M_A \times V$, where M_A is the molarity of A and V is the volume in liters of the solution of A that was used. $5 \text{ mL} = 0.005 \text{ L}$

$$A = 4.0 \text{ M} \times 0.005 \text{ L}$$

$$\underline{0.02} \text{ moles acetone}$$

$$A = 0.02 \text{ moles acetone}$$

- What was the molarity of acetone in the reaction mixture? The volume of the mixture was 50 mL, 0.050 L, and the number of moles of acetone was found in Part (a). Again,

$$M_A = \frac{\text{moles of A}}{\text{Vol soln. in liters}}$$

$$M_A = \frac{0.02 \text{ M}}{0.050 \text{ L}} = 0.4 \text{ M acetone}$$

$$\underline{0.4} \text{ M acetone}$$

- How could you double the molarity of the acetone in the reaction mixture, keeping the total volume at 50 mL and keeping the same concentrations of H^+ ion and I_2 as in the original mixture?

1. Double the moles of acetone concentration

- Using the reaction mixture in Problem 1, a student found that it took 510 seconds for the color of the I_2 to disappear.

- What was the rate of the reaction? Hint: First find the initial concentration of I_2 in the reaction mixture, $[I_2]_0$. Then use Equation 3.

$$[I_2]_0 = 0.0050 \text{ M} \times 0.01 = \frac{0.000050 \text{ mol}}{0.050 \text{ L}} = 0.001 \text{ M}$$

$$\text{rate} = \underline{2.0 \times 10^{-6}} \text{ M/sec}$$

$$\text{rate} = \frac{0.001 \text{ M}}{510 \text{ s}} = 1.96 \times 10^{-6} \text{ M.s}^{-1} \approx 2.0 \times 10^{-6} \text{ M.s}^{-1}$$

- Given the rate from Part (a), and the initial concentrations of acetone, H^+ ion, and I_2 in the reaction mixture, write Equation 3 as it would apply to the mixture.

$$\text{rate} = k [\text{acetone}]^m [I_2]^n [H^+]^p$$

$$2.0 \times 10^{-6} = k (0.4 \text{ M})^m (0.02)^n (0.001)^p$$

- What are the unknowns that remain in the equation in Part (b)?

$$\underline{k} \quad \underline{m} \quad \underline{n} \quad \underline{p}$$

(continued on following page)

$$\begin{aligned} \text{mol } H^+ &= 1.0 \text{ M} \times 0.010 \\ &= \frac{0.010}{0.050} = 0.20 \end{aligned}$$

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Brendan G. Carr



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laboratory All experiments have been thoroughly student tested and include step by step instructions including safety and disposal methods for each

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An Experiment in the Supervision of American History George Gould,1930

Advanced Research on Computers in Education Robert Lewis, Setsuko Ôtsuki,1991 In the field of education advanced research using artificial intelligence and cognitive models is vividly making its appearance clear and deepening our conviction of its effectiveness in practical use This volume is one of the results of the conference which was held with the aim of exploring the current state and future trends of advanced techniques and methodologies in information processing specifically those from artificial intelligence and cognitive science as applied to the field of education

Barron's how to Prepare for the AP Psychology Advanced Placement Examination Robert McEntarffer, Allyson Weseley,2000 This brand new manual offers a diagnostic test and two full length model exams with all questions answered and explained An overview familiarizes test takers with the exam s two parts multiple choice and essay sections It also explains scoring and offers test taking tips An extensive subject review section covers the history of psychology as a scientific discipline the biological basis of behavior sensation and perception states of consciousness learning cognition motivation and emotion developmental psychology personality traits and testing abnormal psychology and its treatments and social psychology Additional material includes extra multiple choice practice questions and an analysis of the essay question with a sample essay and related questions

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