

Experiment 3

Advance Study Assignment: Resolution of Matter into Pure Substances, II. Fractional Crystallization

1. Using Figure 3.1, determine

- a. the number of grams of KNO_3 that will dissolve in 100 g of H_2O at 100°C .

240 g KNO_3

- b. the number of grams of water required to dissolve 25 g of KNO_3 at 100°C . (Hint: your answer to Part a gives you the needed conversion factor for g KNO_3 to g H_2O .)

10.41 g H_2O

- c. the number of grams of water required to dissolve 2.0 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ at 100°C .

1.11 g H_2O

- d. the number of grams of water required to dissolve a mixture containing 25 g KNO_3 and 2.0 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$, assuming that the solubility of one substance is not affected by the presence of another.

11.52 g H_2O

2. To the solution in Problem 1d at 100°C , 10 g of water are added, and the solution is cooled to 0°C .

- a. How much KNO_3 remains in solution?

1 g KNO_3

- b. How much KNO_3 crystallizes out?

24 g KNO_3

- c. How much $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ crystallizes out?

2 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$

- d. What per cent of the KNO_3 in the sample is recovered?

4 %

Advance Study Assignment Experiment 10 Answers

BM King



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1. Understanding the eBook Advance Study Assignment Experiment 10 Answers
 - The Rise of Digital Reading Advance Study Assignment Experiment 10 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Advance Study Assignment Experiment 10 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Advance Study Assignment Experiment 10 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Advance Study Assignment Experiment 10 Answers
 - Personalized Recommendations
 - Advance Study Assignment Experiment 10 Answers User Reviews and Ratings
 - Advance Study Assignment Experiment 10 Answers and Bestseller Lists
5. Accessing Advance Study Assignment Experiment 10 Answers Free and Paid eBooks
 - Advance Study Assignment Experiment 10 Answers Public Domain eBooks
 - Advance Study Assignment Experiment 10 Answers eBook Subscription Services
 - Advance Study Assignment Experiment 10 Answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Advance Study Assignment Experiment 10 Answers Compatibility with Devices
 - Advance Study Assignment Experiment 10 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Advance Study Assignment Experiment 10 Answers
 - Highlighting and Note-Taking Advance Study Assignment Experiment 10 Answers
 - Interactive Elements Advance Study Assignment Experiment 10 Answers
8. Staying Engaged with Advance Study Assignment Experiment 10 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Advance Study Assignment Experiment 10 Answers
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 - Creating a Diverse Reading Collection Advance Study Assignment Experiment 10 Answers
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Advance Study Assignment Experiment 10 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Advance Study Assignment Experiment 10 Answers
 - Fact-Checking eBook Content of Advance Study Assignment Experiment 10 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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