

Unit 1 Atomic Structure & Properties

KEY Terms

Atomic Mass: The mass of an element

Atomic Mass: The mass of a mole worth of an element

Atomic Percent: The percent in mass of something in the overall

Avogadro's #: $6.02 \cdot 10^{23}$

Isotopes: Different types of an element - different number of neutrons

Empirical formula: Reduced version of a formula

Atomic Mass

Where does it come from?

Each element has at least one isotope, and each isotope has a different weight, to find the atomic mass, scientists averaged the weight of each isotope of an element based on how much of each isotope exists in nature.

But what is an Isotope?

Isotopes are different forms of an element and each changes has at least one.

An element is designated based on the number of protons. Electrons determine how exactly an element is

The number of neutrons which makes it is

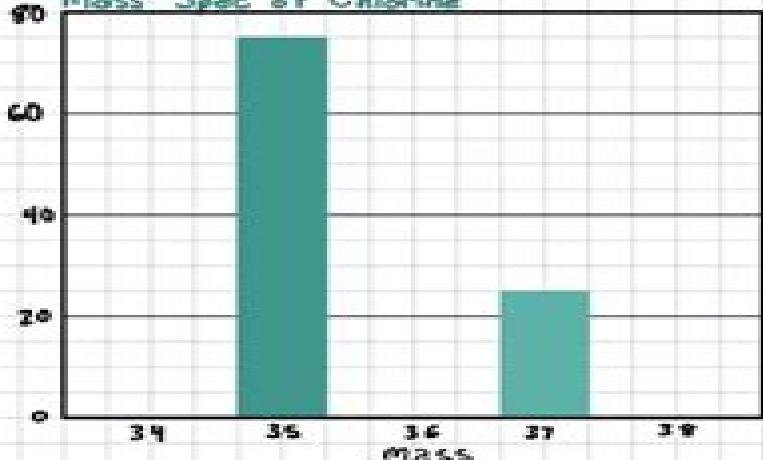
Mass Spectroscopy on the AP Exam

Estimate atomic mass from a mass spectrum graph (could be a fake element)

Identify an element from a mass spectrum graph

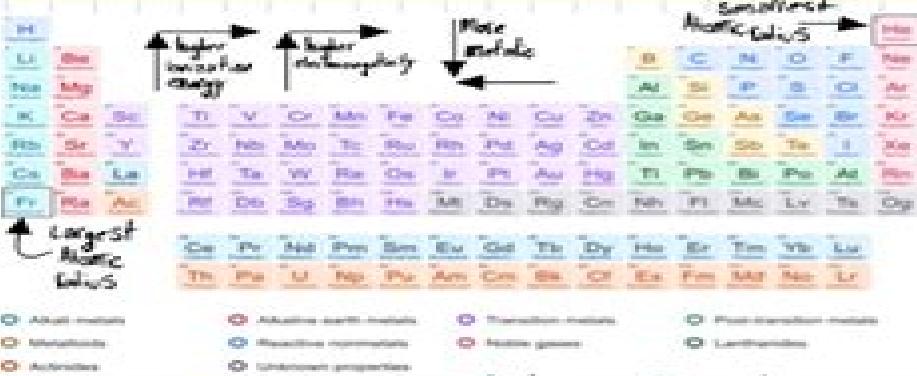
Identify an isotope from a mass spectrum graph

Mass Spec of Chlorine



PERIODIC TABLE OF THE ELEMENTS

1 H	2 He	3 Li	4 Be	5 B	6 C	7 N	8 O	9 F	10 Ne	11 Na	12 Mg	13 Al	14 Si	15 P	16 S	17 Cl	18 Ar	19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr	37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Os	47 Ru	48 Os	49 Ru	50 Os	51 Ru	52 Os	53 Ru	54 Os	55 Ru	56 Os	57 Ru	58 Os	59 Ru	60 Os	61 Ru	62 Os	63 Ru	64 Os	65 Ru	66 Os	67 Ru	68 Os	69 Ru	70 Os	71 Ru	72 Os	73 Ru	74 Os	75 Ru	76 Os	77 Ru	78 Os	79 Ru	80 Os	81 Ru	82 Os	83 Ru	84 Os	85 Ru	86 Os	87 Ru	88 Os	89 Ru	90 Os	91 Ru	92 Os	93 Ru	94 Os	95 Ru	96 Os	97 Ru	98 Os	99 Ru	100 Os	101 Ru	102 Os	103 Ru	104 Os	105 Ru	106 Os	107 Ru	108 Os	109 Ru	110 Os	111 Ru	112 Os	113 Ru	114 Os	115 Ru	116 Os	117 Ru	118 Os
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Approx. Avg. Atomic Mass

$\approx 75\%$ of the isotopes found have a mass of 35 grams

$\approx 25\%$ of the isotopes found have a mass of 37 grams

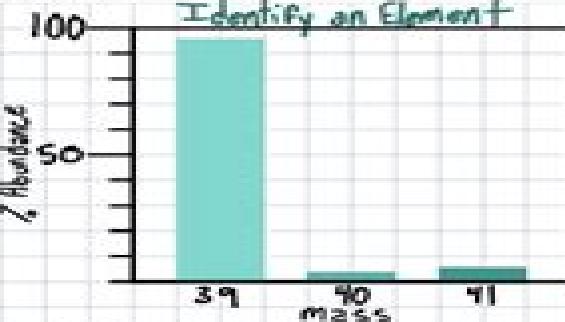
$$\frac{75\%}{100\%} = 0.75 \quad \frac{25\%}{100\%} = 0.25$$

$$(35 \text{ g} \cdot 0.75) + (37 \text{ g} \cdot 0.25)$$

$$26.25 \text{ g} + 9.25 \text{ g}$$

$$35.5 \text{ g}$$

$$\text{avg. g} = \frac{35.5 \text{ g}}{2 \text{ g}}$$



Estimate the atomic mass because most of the mass is 39 grams the element needs to be about 39 grams

1 H	2 He	3 Li	4 Be	5 B	6 C	7 N	8 O	9 F	10 Ne	11 Na	12 Mg	13 Al	14 Si	15 P	16 S	17 Cl	18 Ar	19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr	37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Os	47 Ru	48 Os	49 Ru	50 Os	51 Ru	52 Os	53 Ru	54 Os	55 Ru	56 Os	57 Ru	58 Os	59 Ru	60 Os	61 Ru	62 Os	63 Ru	64 Os	65 Ru	66 Os	67 Ru	68 Os	69 Ru	70 Os	71 Ru	72 Os	73 Ru	74 Os	75 Ru	76 Os	77 Ru	78 Os	79 Ru	80 Os	81 Ru	82 Os	83 Ru	84 Os	85 Ru	86 Os	87 Ru	88 Os	89 Ru	90 Os	91 Ru	92 Os	93 Ru	94 Os	95 Ru	96 Os	97 Ru	98 Os	99 Ru	100 Os	101 Ru	102 Os	103 Ru	104 Os	105 Ru	106 Os	107 Ru	108 Os	109 Ru	110 Os	111 Ru	112 Os	113 Ru	114 Os	115 Ru	116 Os	117 Ru	118 Os
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Potassium makes the most sense as the closest to 39 grams

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William Isaac Fletcher (1844-1917, ed)

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African American Advanced Placement Chemistry Students and Their Developing Study Habits Natalie D. Rasmussen,2006 **Learning and Understanding** National Research Council,Division of Behavioral and Social Sciences and Education,Center for Education,Committee on Programs for Advanced Study of Mathematics and Science in American High Schools,2002-08-06 This book takes a fresh look at programs for advanced studies for high school students in the United States with a particular focus on the Advanced Placement and the International Baccalaureate programs and asks how advanced studies can be significantly improved in general It also examines two of the core issues surrounding these programs they can have a profound impact on other components of the education system and participation in the programs has become key to admission at selective institutions of higher education By looking at what could enhance the quality of high school advanced study programs as well as what precedes and comes after these programs this report provides teachers parents curriculum developers administrators college science and mathematics faculty and the educational research community with a detailed assessment that can be used to guide change within advanced study programs **Library Notes**

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