

AP Biology Study Guide Unit 1

Properties of Water:

1. Polarity - allows H₂O to form Hydrogen Bonds
2. Cohesion - Water H bonds to itself, creating surface tension
3. Adhesion - Water adheres to other polar substances
 - a. Allows easy transport of water against Gravity
4. Universal Solvent - Water can dissolve Polar and Ionic Compounds
 - a. This allows for reactions between Aqueous Solutions
5. Expands when Frozen - this means that Ice floats
6. High Specific Heat - water can absorb or release a large amount of heat with only a slight change in its own temperature (caused by cohesion)
 - a. High Heat of Vaporization - only warmest molecules can evaporate, so this causes evaporative cooling
7. Can be regulated by non-polar hydrophobic substances
8. Neutral PH → can form acids or bases if substances are added

Polarity - Difference in electronegativity between atoms bonded together create slightly charged ends of the molecules they form

Hydrophobic Substance - Nonpolar substance that repels water

Hydrophilic Substance - Polar substance that attracts water molecules

CHON - Carbon, Hydrogen, Nitrogen, Oxygen

- Makeup 99% of living matter
- Strong Biological Cornerstones Because:
 - Smallest Diameter/Valency for atoms of 1, 2, 3, 4 valence electrons
 - Form Stable Covalent Bonds
 - Can form double and triple bonds
 - Versatile

Biological Transition Metals - Mn, Fe, Co, Cu, Zn

- Can exist in multiple oxidation states → can be both oxidized and reduced

Biological Ions: Na⁺, Mg²⁺, Cl⁻, K⁺, Ca²⁺

- Used to maintain homeostasis and osmotic concentration

Phosphorous/Sulphur - Energy Carriers which form strong high energy bonds

5 Foundation Biological Molecules:

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I. Edward Alcamo

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Table of Contents Ap Biology Test Study Guide

1. Understanding the eBook Ap Biology Test Study Guide
 - The Rise of Digital Reading Ap Biology Test Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Biology Test Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Biology Test Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Biology Test Study Guide
 - Personalized Recommendations
 - Ap Biology Test Study Guide User Reviews and Ratings
 - Ap Biology Test Study Guide and Bestseller Lists

5. Accessing Ap Biology Test Study Guide Free and Paid eBooks
 - Ap Biology Test Study Guide Public Domain eBooks
 - Ap Biology Test Study Guide eBook Subscription Services
 - Ap Biology Test Study Guide Budget-Friendly Options
6. Navigating Ap Biology Test Study Guide eBook Formats
 - ePUB, PDF, MOBI, and More
 - Ap Biology Test Study Guide Compatibility with Devices
 - Ap Biology Test Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ap Biology Test Study Guide
 - Highlighting and Note-Taking Ap Biology Test Study Guide
 - Interactive Elements Ap Biology Test Study Guide
8. Staying Engaged with Ap Biology Test Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ap Biology Test Study Guide
9. Balancing eBooks and Physical Books Ap Biology Test Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ap Biology Test Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ap Biology Test Study Guide
 - Setting Reading Goals Ap Biology Test Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Biology Test Study Guide
 - Fact-Checking eBook Content of Ap Biology Test Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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