

Thermochemistry Worksheet Answers

Useful Values:

Heat of Fusion for Water: $\Delta H_{fus} = 6 \text{ kJ/mol}$

Common error: not converting to moles

Heat of Vaporization for Water: $\Delta H_{vap} = 45 \text{ kJ/mol}$

Common error: same

1 calorie = 4.18 Joules

Material	Specific Heat in Joules/g/K	in Cal/g/K	
Water (ice)	2.05	0.490	
Water (liquid)	4.18	1.00	Common error here: using moles instead of grams.
Water (gas)	2.08	0.498	

1. How much energy does it take to warm the following substances the listed amount?

a. 20g of ice from -45° C to -15° C ? $294 \text{ calories / 1230 Joules}$
b. 20g of water from 25° C to 55° C ? $600 \text{ calories / 2508 Joules}$
c. 20g of water vapor from 130° C to 160° C ? $299 \text{ calories / 1249 Joules}$

2. To melt ice, one must bring ice to the freezing point (0° C), and then thaw it into water. How much heat would it take to melt...

a. 30g of ice at 0° C ? 10 kJ $(10,000 \text{ J})$
b. 30g of ice at -40° C ? 12.46 kJ $(12,460 \text{ J})$

3. To boil water, one must raise it to the boiling point (100° C), and then boil it into steam. How much heat would it take to boil...

a. 40g of water at 100° C ? 100 kJ $(100,000 \text{ Joules})$
b. 40g of water at 90° C ? 101.7 kJ $(101,672 \text{ Joules})$
c. 40g of water at 0° C ? 116.7 kJ $(116,720 \text{ Joules})$
d. 40g of ice at 0° C ? 130.1 kJ $(130,053 \text{ Joules})$
e. 40g of ice at -80° C ? 136.6 kJ $(136,613 \text{ Joules})$

4. If 700 calories of heat raise 400g of Phonium by 20° C , what is the specific heat of Phonium? 0.0075 cal/g/K

5. How much heat energy is stored in 2g of ice at -30° C ? (remember: use Kelvin) 2.38 cal / 996 J

6. How much heat energy is stored in 2g of Phonium at 70° C ? (Assume Phonium is a solid at 70° C) 1.672 cal

7. If 100g of water at 10° C comes in contact with _____, what will the final temperature be?
(Remember, you have the specific heat of Phonium from question 4.)

a. 100g of water at 90° C 50° C
b. 100g of Phonium at 90° C (Hint: the water and Phonium end at the same temperature) 16.44° C
Common error: 6.44 degrees is the temperature change, and the water gets warmer.
c. 20g of Phonium at 90° C 11.38° C
d. 1g of Phonium at 90° C 10.67° C

8. If 450g of Phakelite at 100° C is added to 1000g of water at 20° C , the temperature of the water increases by 5° . What is the specific heat of Phakelite? 0.148 cal/g/K

9. If 450g of Nonsenseum at 500° C is added to 1000g of water at 20° C , the temperature of the water increases by 25° . What is the specific heat of Nonsenseum? 0.122 cal/g/K

Answers To Thermochemistry Problems Worksheet

Number Two

**Bull, Prince Hycy, Patterson, Gerrelyn
Chunn**

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