ACID-BASE TITRATION

Name _____

An Acid-Base Titration is the determination of the concentration of an acid or base by exactly neutralizing the acid or base with an acid or base of known concentration.

Example: It takes 45.2 mL of 0.1 M hydrochloric acid to neutralize 30 mL of sodium hydroxide. What was the concentration of sodium hydroxide?

$$NaOH + HCl \Rightarrow NaCl + H_2O$$

Step 1: Express the known concentration in liters .045 L of HCl
Step 2: Multiply by it concentration: (.045)(0.1 M) = .0045
Step 3: Multiply by the mole to mole comparison (see equation)
(.0045) (1/1) = .0045 still
Step 4: Divide by the unknown concentration in liters
.0045/.030 L = .150 M

1. In the titration of 10 mL of vinegar 40 mL of .2 M sodium hydroxide was used. What was the molarity of the vinegar?

$$NaOH + CH_3COOH \Rightarrow H_2O + CH_3COONa$$

_____

2. If 40 mL of .1 M sodium hydroxide neutralizes 80 mL of HCl, what was the molarity of the HCl?

$$NaOH + HCl \Rightarrow NaCl + H_2O$$

_____

3. It took 13 mL at .2 molarity to titrate 45 mL of Ca(OH)$_2$. What was the molarity of the Ca(OH)$_2$?

$$Ca(OH)_2 + 2HCl \Rightarrow CaCl_2 + 2H_2O$$

_____

4. 60 mL of .5 m of HCl is needed to neutralize 60 mL of NaOH. Calculate the molarity of the NaOH

$$NaOH + HCl \Rightarrow H_2O + NaCl$$

_____

5. To titrate 30 mL of potassium hydroxide required 20 mL of .1 M Hydrochloric Acid. Write the equation and tell the molarity of the potassium hydroxide.

_____

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Linda R. Van Thiel

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1. Understanding the eBook Acid Base Titration Virtual Lab Worksheet Answers
   - The Rise of Digital Reading Acid Base Titration Virtual Lab Worksheet Answers
   - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Titration Virtual Lab Worksheet Answers
   - Exploring Different Genres
   - Considering Fiction vs. Non-Fiction
   - Determining Your Reading Goals
3. Choosing the Right eBook Platform
   - Popular eBook Platforms
   - Features to Look for in an Acid Base Titration Virtual Lab Worksheet Answers
   - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Titration Virtual Lab Worksheet Answers

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