

SL Paper 1

The enthalpy of combustion of ethanol is determined by heating a known mass of tap water in a glass beaker with a flame of burning ethanol.

Which will lead to the greatest error in the final result?

- A. Assuming the density of tap water is 1.0 g cm^{-3}
- B. Assuming all the energy from the combustion will heat the water
- C. Assuming the specific heat capacity of the tap water is $4.18 \text{ J g}^{-1} \text{ K}^{-1}$
- D. Assuming the specific heat capacity of the beaker is negligible

What is the enthalpy of combustion of butane in kJ mol^{-1} ?



- A. $4x + 5y - z$
- B. $4x + 5y + z$
- C. $8x + 10y - 2z$
- D. $8x + 5y + 2z$

A 5.00 g sample of a substance was heated from 25.0°C to 35.0°C using $2.00 \times 10^2 \text{ J}$ of energy. What is the specific heat capacity of the substance in $\text{J g}^{-1}\text{K}^{-1}$?

- A. 4.00×10^{-3}
- B. 2.50×10^{-1}
- C. 2.00
- D. 4.00

Which statement is correct for this reaction?



- A. 13.3 kJ are released for every mole of Fe produced.
- B. 26.6 kJ are absorbed for every mole of Fe produced.

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Table of Contents 20math Sl Paper1

1. Understanding the eBook 20math Sl Paper1
 - The Rise of Digital Reading 20math Sl Paper1
 - Advantages of eBooks Over Traditional Books
2. Identifying 20math Sl Paper1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 20math Sl Paper1
 - User-Friendly Interface
4. Exploring eBook Recommendations from 20math Sl Paper1
 - Personalized Recommendations

- 20math Sl Paper1 User Reviews and Ratings
- 20math Sl Paper1 and Bestseller Lists
- 5. Accessing 20math Sl Paper1 Free and Paid eBooks
 - 20math Sl Paper1 Public Domain eBooks
 - 20math Sl Paper1 eBook Subscription Services
 - 20math Sl Paper1 Budget-Friendly Options
- 6. Navigating 20math Sl Paper1 eBook Formats
 - ePub, PDF, MOBI, and More
 - 20math Sl Paper1 Compatibility with Devices
 - 20math Sl Paper1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 20math Sl Paper1
 - Highlighting and Note-Taking 20math Sl Paper1
 - Interactive Elements 20math Sl Paper1
- 8. Staying Engaged with 20math Sl Paper1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 20math Sl Paper1
- 9. Balancing eBooks and Physical Books 20math Sl Paper1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 20math Sl Paper1
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 20math Sl Paper1
 - Setting Reading Goals 20math Sl Paper1
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 20math Sl Paper1
 - Fact-Checking eBook Content of 20math Sl Paper1

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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