

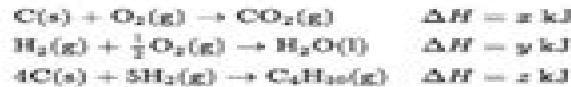
SL Paper 1

The enthalpy of combustion of ethanol is determined by heating a known mass of tap water in a glass beaker with a flame of burning ethanol.

Which will lead to the greatest error in the final result?

- A. Assuming the density of tap water is 1.0 g cm^{-3}
- B. Assuming all the energy from the combustion will heat the water
- C. Assuming the specific heat capacity of the tap water is $4.18 \text{ J g}^{-1} \text{ K}^{-1}$
- D. Assuming the specific heat capacity of the beaker is negligible

What is the enthalpy of combustion of butane in kJ mol^{-1} ?



- A. $4x + 8y - z$
- B. $4x + 8y + z$
- C. $8x + 10y - 2z$
- D. $8x + 8y + 2z$

A 5.00 g sample of a substance was heated from 25.0°C to 35.0°C using $2.00 \times 10^2 \text{ J}$ of energy. What is the specific heat capacity of the substance

in $\text{J g}^{-1} \text{ K}^{-1}$?

- A. 4.00×10^{-3}
- B. 2.50×10^{-1}
- C. 2.00
- D. 4.00

Which statement is correct for this reaction?



- A. 13.3 kJ are released for every mole of Fe produced.
- B. 26.6 kJ are absorbed for every mole of Fe produced.

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