

KEY

Electron Review Questions

Using the equations, solve the following:

$$c = \lambda\nu \quad c = \text{speed of light } 3.0 \times 10^8 \text{ m/s} \quad \nu = \text{frequency} \quad \lambda = \text{wavelength}$$

$$E = h\nu \quad E = \text{energy in Joules} \quad h = \text{Planck constant, } 6.626 \times 10^{-34} \quad \nu = \text{frequency}$$

1. What is the frequency of a wave that has a wavelength of 4×10^{-7} meters? $\nu = \frac{c}{\lambda} = \frac{3.0 \times 10^8}{4 \times 10^{-7}} = 7.50 \times 10^{14} \text{ s}^{-1}$

2. What is the energy of this wave? $E = h\nu = 6.626 \times 10^{-34} \times 7.50 \times 10^{14} = 4.95 \times 10^{-19} \text{ J}$

3. List 5 forms of electromagnetic radiation and what is the lowest energy color of light? Which color of light has the highest energy?

Radio waves, microwaves, visible light, UV, X-rays, Gamma

4. What is the relationship between frequency and wavelength?

It is ~~inversely proportional~~, i.e., as wavelength decreases frequency increases.

5. What is the relationship between energy and frequency?

It is ~~directly proportional~~, as frequency increases, energy increases.

6. Where are the electrons located in the atom?

Well outside the nucleus in the electron cloud

7. Define:

a) Ground State - lowest energy state of an atom, ~~atom~~

b) Excited State - state with atoms gain energy

8. What is a bright-line or line-emission spectrum of an atom?

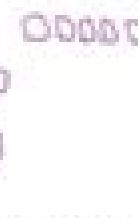
A series of ~~specific~~ frequencies and wavelengths. For the light that ~~emanates~~ when an element loses its ~~energy~~ through a process

9. The principle that says no two electrons can have the same set of quantum numbers is Pauli Exclusion Principle.

10. The principle that says electrons will orbit the lowest energy orbitals first is called Aufbau principle.

11. The rule that says electrons will occupy orbitals of equal energy alone, before pairing up is called Hund's Rule.

12. Draw the Aufbau Diagram through 3d.



13. Name where the blocks located on the periodic table:

- s block: ~~1s + 2 electrons~~
- p block: ~~2s + 6 electrons~~
- d block: ~~3s + 10 electrons, 3p + 6 electrons, 3d + 10 electrons~~
- f block: ~~4s + 14 electrons~~

14. What is the shape of the p orbital? Spherical

~~Orbital~~, spherical

15. Write the electron configuration using orbital notation for sulfur:



16. Write the electron configuration for Cadmium using electron configuration notation:



17. Write the electron configuration for lead using noble gas notation:



18. How many valence electrons does sulfur have?

~~6~~, 6

19. Why are copper and chromium exceptions to the Aufbau principle?

~~They break the 1s from 4s and holding 3d for Cr and anything after 3d for Cu.~~

20. Write the electron configuration for the element Sb using the Noble gas abbreviation:



Answer Key Unit 3 Review Electrons

Kenneth R. Miller

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