

KEV

Electron Review Questions

Using the equations, solve the following:

$$c = \lambda \nu \quad c = \text{speed of light } 3.0 \times 10^8 \text{ m/sec} \quad \nu = \text{frequency} \quad \lambda = \text{wavelength}$$

$$E = h\nu \quad E = \text{energy in joules} \quad h = \text{planck constant } 6.626 \times 10^{-34} \quad \nu = \text{frequency}$$

1. What is the frequency of a wave that has a wavelength of 4×10^7 meters?
 $\nu = \frac{c}{\lambda} = \frac{3.0 \times 10^8 \text{ m/sec}}{4 \times 10^7 \text{ m}} = 7.50 \times 10^{-14} \text{ s}^{-1}$

2. What is the energy of this wave?

$$E = h\nu = 6.626 \times 10^{-34} \times 7.5 \times 10^{-14} = 4.97 \times 10^{-47} \text{ J}$$

3. List 5 forms of electromagnetic radiation and what is the lowest energy color of light? Which color of light has the highest energy?

Radio waves, microwaves, visible light, UV, x-rays, gamma

4. What is the relationship between frequency and wavelength?

IT is inversely proportional. i.e., as wavelength decreases frequency increases it

5. What is the relationship between energy and frequency?

IT is directly proportional. As frequency increases, energy increases

6. Where are the electrons located in the atom?

well around the nucleus in the electron cloud

7. Define:

a) Ground State - lowest energy state of an atom, a neutral atom

b) Excited State - state when atom gains energy

8. What is a bright-line or line-emission spectrum of an atom?

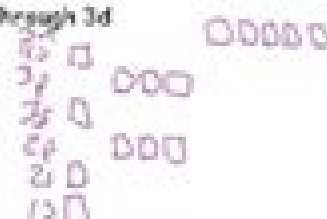
a series of specific frequencies and wavelengths of visible light that appear when an element's light is passed through a prism

9. The principle that says no two electrons can have the same set of quantum numbers is Pauli Exclusion Principle

10. The principle that says electrons will orbit the lowest energy orbitals first is called Aufbau principle

11. The rule that says electrons will occupy orbitals of equal energy alone, before pairing up is called Hund's rule

12. Draw the Aufbau Diagram through 3d



13. Name where the blocks located on the periodic table:

a) s block? 1st & 2nd columns

b) p block? right 6 columns

c) d block? in the middle between s and p (8 gray boxes)

d) f block? two rows at the bottom

14. What is the shape of the p orbital? s orbital?

dumbbell, spherical

15. Write the electron configuration using orbital notation for sulfur:



16. Write the electron configuration for Cadmium using electron configuration notation.



17. Write the electron configuration for lead using noble gas notation:



18. How many valence electrons does sulfur have?



19. Why are copper and chromium exceptions to the Aufbau principle?

they borrow 1e⁻ from 4s and halfing fill 3d for Cr and completely fill 3d for Cu

20. Write the electron configuration for the element Sb using the Noble gas abbreviation.



Answer Key Unit 3 Review Electrons

Kenneth R. Miller



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