

14: Chemical Bonding Theories

Key Chemistry Terms

- **Ionic Bond:** Bond formed from electrostatic attraction between ions (charged atoms)—formed from metals with nonmetals.
- **Covalent Bond:** A bond formed between two nonmetals that involves shared electrons.
- **Polar covalent bond:** Bond formed between two nonmetals—uneven sharing of electrons.
- **Metallic bond:** Formed between metal atoms—electrons are pooled between the network of atoms.
- **Polar bond:** Bond between two atoms with a great difference in electronegativities—uneven sharing of electrons resulting in a slightly positive region and a slightly negative region of the bond.
- **Electronegativity:** An atom's "pull" on electrons shared with another atom.
- **Polar molecules:** Molecule where the polar bonds do not cancel each other out in 3D orientation.
- **Isomers:** Molecules with the same chemical formula, but different bonding structure.
- **Resonance:** Two compounds with multiple bonds in different locations.
- **Valence Bond Theory:** Overlap of atomic orbitals form bond.
- **Sigma (σ) bonds:** First bond between two atoms formed from head on overlap of orbitals.
- **Pi (π) bonds:** 2nd or 3rd bond between two atoms formed from overlap of parallel p orbitals.
- **Hybridization:** Atomic orbitals hybridize into orbitals of the same energy for all sigma bonds of the atom.

Characteristics of bond types

Bond types	Properties	Electron source
ionic	Metals + non-metals	Transfer of e ⁻
Covalent	Non-metals	Shared
Polar	Non-metals	Shared
Coordinate		unusually
Metallic	Metals	delocalised

Comments: a few individuals.

- **Ionics:** High melting points, most dissolve in water, conduct electricity when dissolved in water, brittle.
- **Covalents:** Low melting points, most do not dissolve in water, do not conduct electricity when dissolved in water.
- **Polar covalents:** Medium melting points, some dissolve in water, do not conduct electricity when dissolved in water.
- **Metallic:** Soft, conduct heat and electricity, do not dissolve in water.

Bond Polarity

When nonmetals bond covalently with a large difference in electronegativity,

- Absolute value of differences:
 - $0 - 0.4 = \text{covalent}$
 - $0.5 - 1.4 = \text{polar covalent}$
 - $1.5 - 4 = \text{ionic}$

Symbolized with an arrow pointing towards the more electronegative element and a crossed tail by the less electronegative element.

Example:
C - N Electronegativity difference of 0.5 = polar bond



Isomers and Resonance

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FRANCESCO CROCI, *Università di Padova*



Sigma and Pi bonds

How many sigma and pi bonds are in the following?



6 sigma bonds, 6 pi bonds

Hybridization

# of sigma bonds	Hybridization	Formed from
2	sp	1 s and 1 p
3	sp ²	1 s and 2 p
4	sp ³	1 s and 3 p

How to Use This Cheat Sheet: These are the keys related this topic. Try to read through it carefully twice then rewrite it out on a blank sheet of paper. Review it again before the exams.

Answer Key To The Physical Setting Chemistry

Leveson Francis Vernon-Harcourt



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