

Acids & Bases HOMEWORK

Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrofluoric acid**
- b. **HI**: **Hydroiodic acid**
- c. **HNO₃**: **nitric acid**
- d. **HClO₄**: **perchloric acid**
- e. **HBrO₃**: **hypobromous acid**
- f. **HClO₄**: **chloric acid**
- g. **H₂SO₄**: **hydrochloric acid**
- h. **H₂SO₄**: **hydrobromic acid**
- i. **H₂SO₄**: **hydroiodic acid**
- j. **H₂SO₄**: **hydrochloric acid**

2. Of the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. Write the dissociation for **Na₂CO₃**.

4. **NaClO₄** will react with water and lose a ClO_4^- ion in the process.

a. Write a balanced equation for this reaction: $\text{NaClO}_4 + \text{H}_2\text{O} \longrightarrow \text{Na}^+ + \text{HClO}_4^-$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating HClO_4^- in a weak acid.

c. Dissolve **NaClO₄** in a portion of water and a portion of ethanol? **commonly** dissociating in both a pH **base**.

d. In your **reaction** prediction, is **NaClO₄** acting as a **Brønsted-Lowry acid** or a **Brønsted-Lowry base**? **Brønsted-Lowry acid**. Because it is a portion of a **base**.

e. Identify the conjugate base of **NaClO₄**: **ClO₄⁻**

f. Identify **NaClO₄** as **strong**, **weak**, **intermediate**, or **none**: **strong**, because it is a portion of a **strong** base, **NaOH**.

5. **Li⁺** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{Li}^+ + \text{H}_2\text{O} \longrightarrow \text{H}^+ + \text{LiOH}$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating LiOH in a weak base.

c. Dissolve **LiOH** in a portion of water and a portion of ethanol? **commonly** dissociating in both a pH **base**.

d. In your **reaction** prediction, is **LiOH** acting as a **Brønsted-Lowry acid** or a **Brønsted-Lowry base**? **Brønsted-Lowry base**. Because it is a portion of a **base**.

e. Identify the conjugate acid of **LiOH**: **Li⁺**

f. Identify **LiOH** as **strong**, **weak**, **intermediate**, or **none**: **strong**

6. **NaClO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{NaClO}_4 + \text{H}_2\text{O} \longrightarrow \text{Na}^+ + \text{HClO}_4^-$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating HClO_4^- in a weak acid.

c. Dissolve **NaClO₄** in a portion of water and a portion of ethanol? **commonly** dissociating in both a pH **base**.

d. In your **reaction** prediction, is **NaClO₄** acting as a **Brønsted-Lowry acid** or a **Brønsted-Lowry base**? **Brønsted-Lowry acid**. Because it is a **portion** of a **base**.

e. Identify the conjugate base of **NaClO₄**: **ClO₄⁻**

f. Identify **NaClO₄** as **strong**, **weak**, **intermediate**, or **none**: **strong**, because it is a portion of a **strong** base, **NaOH**. Because it is a **portion** of a **base**.

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Henry Watts, Henry Forster

Morley, Matthew Moncrieff Pattison

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