

Section 4.2 The Structure of an Atom-Physical Science

1. What are 3 subatomic particles? - answer a. Protons
b. Electrons
c. Neutrons
2. Circle the letter that identifies a subatomic particle with a positive (+) charge. - answerb. Proton
3. Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? - answerA charged object did not deflect their path
4. Circle the letters of properties that vary among subatomic particles. - answera. Color
c. Charge
d. Location in the Atom
5. Circle the letter of the expression that accurately compares the masses of neutrons and protons. - answera. Mass of 1 neutron=mass of 1 proton
6. Two atoms of the same element can have different numbers of protons. - answerTrue
7. What is an atomic number? - answerThe number of protons in an atom
8. Circle the letters that identify quantities that are always equal to an element's atomic number. - answerb. Number of Protons
c. Number of Neutrons
d. Number of Electrons
9. Two different elements can have the same atomic number. - answerTrue
10. What is the mass number of an atom? - answerThe sum of the protons and neutrons in an atom
(# of protons + # of neutrons)
11. Complete the equation... - answer# of neutrons=mass # - atomic # (# of protons)
12. Every atom of a given element has the same number of _____ and _____ - answerProtons and Electrons

Answers To Section 2 The Atom Packet

John R. Holum

Answers To Section 2 The Atom Packet:

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