



# ISOTOPES PRACTICE SET



1. Write the symbol for the following isotopes:

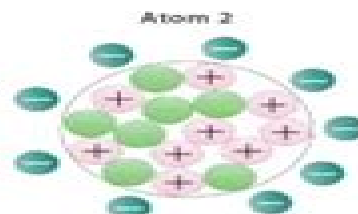
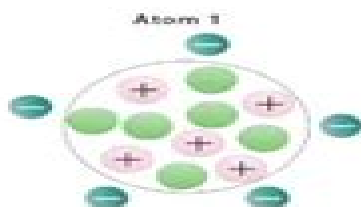
a. 8 protons and 8 neutrons

c. atomic number is 11 and mass number is 23

b. 28 protons and 30 neutrons

d. 92 protons and mass number is 238

2. Fill in the table for the two atoms shown below.



|                    | Atom 1 | Atom 2 |
|--------------------|--------|--------|
| Number of Protons  |        |        |
| Number of Neutrons |        |        |
| Mass Number        |        |        |

3. Fill in the following table.

|                    |                        |                     |    |
|--------------------|------------------------|---------------------|----|
| Isotope Symbol     | ${}^{40}_{19}\text{K}$ | ${}^{18}_9\text{F}$ |    |
| Atomic Number      |                        |                     | 16 |
| Mass Number        |                        |                     |    |
| Number of Protons  |                        |                     |    |
| Number of Neutrons |                        |                     | 15 |

Name : \_\_\_\_\_

Date : \_\_\_\_\_

# Answers To Isotope Practice Chemistry 2 Points

**Steven W. Dulan, Amy Dulan**



## Answers To Isotope Practice Chemistry 2 Points:

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