

## Half-Life

Textbook pages 302–311

### Before You Read

Write a sentence in the lines below explaining what the word decay means to you. As you read about radioactive decay, think about how the common meaning of decay differs from the scientific meaning.

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### What is radiocarbon dating?

**Radiocarbon dating** is the process of determining the age of an object by measuring the amount of carbon-14 remaining in that object. Carbon's isotopes include carbon-12 and carbon-14. When an organism is alive, the ratio of carbon-14 atoms to carbon-12 atoms in the organism remains nearly constant. But when an organism dies, its carbon-14 atoms decay without being replaced. The ratio of carbon-14 to carbon-12 then decreases with time. By measuring this ratio, the age of an organism's remains can be estimated. Only material from plants and animals that lived within the past 50 000 years contains enough carbon-14 to be measured using radiocarbon dating. ☐

### What is a half-life and how is it used in radiocarbon dating?

A half-life is a measure of the rate of radioactive decay for a given isotope. It is equal to the time required for half the nuclei in a sample to decay. Its value is a constant for any radioactive isotope. For example, the half-life of the radioisotope strontium-90 is 29 years. If you have 10.0 g of strontium-90 today, 29 years from now you will have 5.00 g left. This is because one half-life will have passed ( $10.0 \text{ g} \times \frac{1}{2} = 5.00 \text{ g}$ ). 58 years from now, two half-lives will have passed and 2.50 g of the sample will remain ( $10.0 \text{ g} \times \frac{1}{2} \times \frac{1}{2} = 2.50 \text{ g}$ ). The shorter the half-life is, the faster the decay rate. A **decay curve** is a curved line on a graph that shows the rate at which radioisotopes decay.

### Mark the Text

#### In Your Own Words

After you read this section, go back and summarize the main concepts in your own words.

### Reading Check

1. Which carbon isotope undergoes radioactive decay?

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## Answers To Half Life 7 2 Summary

**Ralph H. Petrucci, Robert K. Wissner**

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