

# *The Banker Babe*

by Argus



Modern Erotic Library

# The Banker Babe Modern Erotic Library English Edition

**DJ Losen**

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 6\text{e}^- \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$  make electrons equal convert to basic solution  $4\text{Cr}_2\text{O}_7^{2-} + 14\text{H}_2\text{O} + 6\text{e}^- \rightarrow 8\text{Cr}^{3+} + 7\text{H}_2\text{O}_2$  the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce  $\text{H}_2\text{O}$

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add  $\text{OH}^-$  ions or the  $\text{OH}^-$  pair to fully balance the equation in this video we'll walk through this process for the reaction between  $\text{ClO}_4^-$  and  $\text{Cr(OH)}_4^-$  in basic solution created by jay

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni}^{2+} \rightarrow \text{Ni}^{3+}$  2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add h ions or the h h<sub>2</sub>o pair to fully balance the equation in this video we'll walk through this process for the reaction between dichromate  $\text{Cr}_2\text{O}_7^{2-}$  and chloride cl ions in acidic solution created by jay

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