

Experiment 3

Advance Study Assignment: Resolution of Matter into Pure Substances. II. Fractional Crystallization

1. Using Figure 3.1, determine

- a. the number of grams of KNO_3 that will dissolve in 100 g of H_2O at 100°C .

240 g KNO_3

- b. the number of grams of water required to dissolve 25 g of KNO_3 at 100°C . (Hint: your answer to Part a gives you the needed conversion factor for g KNO_3 to g H_2O .)

10.41 g H_2O

- c. the number of grams of water required to dissolve 2.0 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ at 100°C .

1.11 g H_2O

- d. the number of grams of water required to dissolve a mixture containing 25 g KNO_3 and 2.0 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$, assuming that the solubility of one substance is not affected by the presence of another.

11.52 g H_2O

2. To the solution in Problem 1d at 100°C , 10 g of water are added, and the solution is cooled to 0°C .

- a. How much KNO_3 remains in solution?

1 g KNO_3

- b. How much KNO_3 crystallizes out?

24 g KNO_3

- c. How much $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ crystallizes out?

2 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$

- d. What per cent of the KNO_3 in the sample is recovered?

4 %

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