

SECTION E

The following questions are long answer type and carry 5marks each. Two questions have an internal choice.

33. i) The value of K_c for the reaction

$3\text{O}_2(\text{g}) \rightleftharpoons 2\text{O}_3(\text{g})$ is 2.0×10^{-5} at 25°C . If the equilibrium concentration of O_2 in air at 25°C is 1.6×10^{-2} . What is the concentration of O_3 .

ii) The species HCO_3^- and NH_3 can act as Bronsted acid and bases. For each case give the corresponding conjugate acid and base.

iii) If $Q_c < K_c$, in which direction reaction will proceed?

34. a) Calculate the heat of combustion of glucose from the following data:



b) Define the following terms:

i) Enthalpy of neutralization

ii) Hess's law of constant heat summation.

OR

a) Calculate the temperature above which the reduction of lead oxide in the following reaction becomes spontaneous.



[Given $\Delta H = 108.4 \text{ kJ/mol}$, $\Delta S = 190 \text{ J K}^{-1} \text{ mol}^{-1}$]

b) Define Entropy and State Function.

35. a) Explain the mechanism of nitration in benzene ring.

b) Arrange Benzene, n - Hexane and Ethyne in decreasing order of acidic behaviour. Also give reason for this behaviour

OR

a) Explain the mechanism of Friedel Craft Alkylation reaction.

b) Why is Wurtz reaction not preferred for the alkanes containing odd number of carbon atoms?

c) How will you convert benzene into acetophenone.

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Eric John Holmyard

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