

Mole Conversions Worksheet #1

1. Mole \rightarrow Mass Conversions – using molar mass of each substance, convert the following quantities.

- | | |
|--|---|
| a. 10.0 mol Cr 520 g | f. 0.160 mol H ₂ O 2.88 g |
| b. 3.32 mol K 130 g | g. 5.08 mol Ca(NO ₃) ₂ 834 g |
| c. 2.20×10^{-3} mol Sn 0.261 g | h. 15.0 mol H ₂ SO ₄ 1470 g |
| d. 0.720 mol Be 6.48 g | i. 4.52×10^{-2} mol C ₂ H ₆ 1.27 \times 10^{-3} g |
| e. 2.40 mol N ₂ 67.2 g | j. 0.0112 mol K ₂ CO ₃ 1.55 g |

2. Mass \rightarrow Mole Conversions – using molar mass of each substance convert the following quantities.

- | | |
|--|---|
| a. 72.0 g Ar 1.80 mol | f. 27.4 g NO ₂ 0.996 mol |
| b. 3.70×10^{-2} g B 3.43 \times 10^{-2} mol | g. 5.00 g H ₂ 2.50 mol |
| c. 187 g Al 6.93 mol | h. 2.64×10^{-4} g Li ₂ PO ₄ 2.28 \times 10^{-4} mol |
| d. 3.33 g SnF ₂ 2.13 mol | i. 11.0 g CH ₄ 0.688 mol |
| e. 7.21×10^{-2} g He 1.80 \times 10^{-2} mol | j. 847 g (NH ₄) ₂ CO ₃ 8.82 mol |

3. What is the volume of the following gases?

- | |
|---|
| a. 5.40 mol O ₂ 121 L |
| b. 3.20×10^{-2} mol CO ₂ 0.717 L |
| c. 0.960 mol SO ₂ 21.5 L |

4. How many moles are in each of the following volumes?

- | |
|---|
| a. 89.6 L Ne 4.00 mol |
| b. 1.00×10^3 L C ₂ H ₆ 44.6 mol |
| c. 5.42×10^4 F ₂ 2.42 \times 10^4 mol |

5. Find the number of moles in each of the number of representative particles.

- | |
|---|
| a. 1.20×10^{23} atoms of P 19.9 mol |
| b. 3.87×10^{23} molecules of AlF ₃ 6.43 \times 10^{23} mol |
| c. 4.81×10^{24} molecules of NH ₃ 7.99 \times 10^{24} mol |

6. How many representative particles are in each of the following mole quantities?

- | |
|--|
| a. 1.24 mol Cl ₂ 7.46 \times 10^{23} molecules |
| b. 4.20×10^{-2} mol K ₂ S 2.53 \times 10^{22} molecules |
| c. 34.02 mol Ca(OH) ₂ 2.048 \times 10^{25} molecules |

7. Convert the following two-step quantities, converting first to moles and then to the desired quantity.

- | |
|--|
| a. Find the number of molecules in 60.0 g of N ₂ O. 8.21×10^{23} molecules |
| b. Find the volume of 3.24×10^{22} molecules of Ne. 1.21 L |
| c. Find the mass of 18.0 L of CH ₄ . 12.9 g |
| d. Find the volume of 835 g of SO ₂ . 234 L |
| e. Find the mass of one atom of nickel. 1×10^{-27} g |

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