

# Mole Conversions Worksheet #1

1. Mole  $\rightarrow$  Mass Conversions – using molar mass of each substance, convert the following quantities.

- |  |   |
|--|---|
| a. 10.0 mol Cr <b>520 g</b>                    | f. 0.160 mol $\text{H}_2\text{O}$ <b>2.88 g</b>   |
| b. 3.32 mol K <b>130 g</b>                     | g. 5.08 mol $\text{Ca}(\text{NO}_3)_2$ <b>834 g</b>   |
| c. $2.20 \times 10^{-5}$ mol Sn <b>0.261 g</b> | h. 15.0 mol $\text{H}_2\text{SO}_4$ <b>1470 g</b>   |
| d. 0.720 mol Be <b>6.48 g</b>                  | i. $4.52 \times 10^{-5}$ mol $\text{C}_2\text{H}_6$ <b><math>1.27 \times 10^{-3}</math> g</b> |
| e. 2.40 mol $\text{N}_2$ <b>67.2 g</b>         | j. 0.0112 mol $\text{K}_2\text{CO}_3$ <b>1.55 g</b>   |

2. Mass  $\rightarrow$  Mole Conversions – using molar mass of each substance convert the following quantities.

- |   |   |
|---|---|
| a. 72.0 g Ar <b>1.80 mol</b>  | f. 27.4 g $\text{NO}_2$ <b>0.596 mol</b>  |
| b. $3.70 \times 10^{-5}$ g B <b><math>3.43 \times 10^{-7}</math> mol</b>  | g. 5.00 g $\text{H}_2$ <b>2.50 mol</b>  |
| c. 187 g Al <b>6.93 mol</b>   | h. $2.64 \times 10^{-4}$ g $\text{Li}_3\text{PO}_4$ <b><math>2.28 \times 10^{-6}</math> mol</b> |
| d. 333 g $\text{SnF}_2$ <b>2.13 mol</b>                                   | i. 11.0 g $\text{CH}_4$ <b>0.688 mol</b>  |
| e. $7.21 \times 10^{-5}$ g He <b><math>1.80 \times 10^{-7}</math> mol</b> | j. 847 g $(\text{NH}_4)_2\text{CO}_3$ <b>8.82 mol</b>   |

3. What is the volume of the following gases?

- a. 5.40 mol  $\text{O}_2$  **121 L**  
b.  $3.20 \times 10^{-2}$  mol  $\text{CO}_2$  **0.717 L**  
c. 0.960 mol  $\text{SO}_3$  **21.5 L**

4. How many moles are in each of the following volumes?

- a. 89.6 L Ne **4.00 mol**  
b.  $1.00 \times 10^5$  L  $\text{C}_2\text{H}_6$  **44.6 mol**  
c.  $5.42 \times 10^1$   $\text{F}_2$   **$2.42 \times 10^{-2}$  mol**

5. Find the number of moles in each of the number of representative particles.

- a.  $1.20 \times 10^{25}$  atoms of P **19.9 mol**  
b.  $3.87 \times 10^{23}$  molecules of  $\text{AlF}_3$   **$6.45 \times 10^{-3}$  mol**  
c.  $4.81 \times 10^{24}$  molecules of  $\text{NH}_3$   **$7.99 \times 10^{-3}$  mol**

6. How many representative particles are in each of the following mole quantities?

- a. 1.24 mol  $\text{Cl}_2$   **$7.46 \times 10^{23}$  molecules**  
b.  $4.20 \times 10^{-5}$  mol  $\text{K}_2\text{S}$   **$2.53 \times 10^{23}$  molecules**  
c. 34.02 mol  $\text{Ca}(\text{OH})_2$   **$2.048 \times 10^{25}$  molecules**

7. Convert the following two-step quantities, converting first to moles and then to the desired quantity.

- a. Find the number of molecules in 60.0 g of  $\text{N}_2\text{O}$ .  **$8.21 \times 10^{23}$  molecules**  
b. Find the volume of  $3.24 \times 10^{22}$  molecules of Ne. **1.21 L**  
c. Find the mass of 18.0 L of  $\text{CH}_4$ . **12.9 g**  
d. Find the volume of 835 g of  $\text{SO}_3$ . **234 L**  
e. Find the mass of one atom of nickel.  **$1 \times 10^{-22}$  g**

# Answer Key Mass Mole Conversion Worksheet

**ML Yell**



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