

1. Boyle's Law: When temperature is held constant, the pressure and volume of a gas are inversely proportional.

2. Mathematically, Boyle's law is stated $PV = \frac{k}{P_2V_2}$ or $P_1V_1 = P_2V_2$

3. At a pressure of 405 kPa, the volume of a gas is 6.00 cm³. Assuming the temperature remains constant, at what pressure will the new volume be 4.00 cm³?

$$\begin{aligned} P_1 &= 405 \text{ kPa} \\ V_1 &= 6.00 \text{ cm}^3 \end{aligned}$$

$$\begin{aligned} P_2 &= ? \\ V_2 &= 4.00 \text{ cm}^3 \end{aligned}$$

$$P_1V_1 = P_2V_2$$

$$405 \text{ kPa} \times 6.00 \text{ cm}^3 = P_2 \times 4.00 \text{ cm}^3 \quad P_2 = \boxed{608 \text{ kPa}}$$

4. A volume of gas at 1.10 atm was measured at 326 cm³. What will be the volume if the pressure is adjusted to 1.90 atm?

$$\begin{aligned} P_1 &= 1.10 \text{ atm} \\ V_1 &= 326 \text{ cm}^3 \end{aligned}$$

$$\begin{aligned} P_2 &= 1.90 \text{ atm} \\ V_2 &= ? \end{aligned}$$

$$P_1V_1 = P_2V_2$$

$$1.10 \text{ atm} \times 326 \text{ cm}^3 = 1.90 \text{ atm} \times V_2 \quad V_2 = \boxed{189 \text{ cm}^3}$$

5. If 36.5 m³ of a gas are collected at a pressure of 755 mm Hg, what volume will the gas occupy if the pressure is changed to 632 mm Hg?

$$\begin{aligned} P_1 &= 755 \text{ mm Hg} \\ V_1 &= 36.5 \text{ m}^3 \end{aligned}$$

$$\begin{aligned} P_2 &= 632 \text{ mm Hg} \\ V_2 &= ? \end{aligned}$$

$$P_1V_1 = P_2V_2$$

$$755 \text{ mm Hg} \times 36.5 \text{ m}^3 = 632 \text{ mm Hg} \times V_2 \quad V_2 = \boxed{47.6 \text{ m}^3}$$

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Table of Contents Answer Key For Boyles Law

1. Understanding the eBook Answer Key For Boyles Law
 - The Rise of Digital Reading Answer Key For Boyles Law
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Boyles Law
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Boyles Law
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Boyles Law
 - Personalized Recommendations
 - Answer Key For Boyles Law User Reviews and Ratings
 - Answer Key For Boyles Law and Bestseller Lists

5. Accessing Answer Key For Boyles Law Free and Paid eBooks
 - Answer Key For Boyles Law Public Domain eBooks
 - Answer Key For Boyles Law eBook Subscription Services
 - Answer Key For Boyles Law Budget-Friendly Options
6. Navigating Answer Key For Boyles Law eBook Formats
 - ePUB, PDF, MOBI, and More
 - Answer Key For Boyles Law Compatibility with Devices
 - Answer Key For Boyles Law Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key For Boyles Law
 - Highlighting and Note-Taking Answer Key For Boyles Law
 - Interactive Elements Answer Key For Boyles Law
8. Staying Engaged with Answer Key For Boyles Law
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key For Boyles Law
9. Balancing eBooks and Physical Books Answer Key For Boyles Law
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key For Boyles Law
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key For Boyles Law
 - Setting Reading Goals Answer Key For Boyles Law
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key For Boyles Law
 - Fact-Checking eBook Content of Answer Key For Boyles Law
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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