

Stoichiometry "End of Chapter" IN CLASS Practice Assessment

Background Data -

- Why are the coefficients in a balanced equation necessary for stoichiometry?
Coefficients are necessary for establishing a mole ratio ratio to convert between moles of different substances.

Skills - Using a balanced equation, calculate using mole to mole ratios. Using a balanced equation, given the grams of reactant, calculate the grams of product.

- Consider the following unbalanced equation.



There are 5.00g of manganese(IV) oxide present. calculate the grams of chlorine gas.
 $5.00\text{g MnO}_2 \times 1\text{mol MnO}_2 \times 1\text{mol Cl}_2 \times 70.9\text{g Cl}_2 = 45.45\text{g Cl}_2$
 $55.94\text{g MnO}_2 \cdot 1\text{mol MnO}_2 \cdot 1\text{mol Cl}_2$

Skills - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Since it is incomplete, you calculate that a reaction should produce 4.0g Cu(OH)₂. After the lab is over you find that 3.8g Cu(OH)₂ were collected. What would the percent yield be for that reaction completed within the lab?

$$\text{% yield} = \frac{\text{Actual/Theoretical}}{100} \times 100$$

$$= 0.95/1.00 \times 100 = 95\%$$

Skills - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate reacts with aqueous hydrochloric acid to produce solid calcium chloride, carbon dioxide, and water vapor. What is the percent yield if 5.00g of calcium carbonate reacts with 25.0g of hydrochloric acid and only 3.42g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$1.00\text{g CaCO}_3 \times 1\text{mol CaCO}_3 / 100.0\text{g CaCO}_3 = 0.010\text{mol CaCO}_3 / 1\text{mol in rxn} = 1.00\text{ mol}$$

$$25.0\text{g HCl} \times 1\text{mol HCl} / 73\text{g HCl} = 0.342\text{ mol HCl} / 2\text{mol in rxn} = 0.171\text{ mol}$$

$$0.010\text{mol CaCO}_3 \times (1\text{mol CaCl}_2 / 1\text{mol CaCO}_3) \times (110.9\text{g CaCl}_2 / 1\text{mol CaCl}_2) = 1.11\text{g}$$

$$\text{% yield} = \frac{\text{Actual/Theoretical}}{100} \times 100$$

$$= 1.11\text{g} / 1.71\text{g} \times 100 = 65.55\%$$

Assessment Stoichiometry Answer Key

Nathan Daniel Price

Assessment Stoichiometry Answer Key:

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This book delves into Assessment Stoichiometry Answer Key. Assessment Stoichiometry Answer Key is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Assessment Stoichiometry Answer Key, encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to Assessment Stoichiometry Answer Key
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- Chapter 5: Conclusion

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6. In chapter 5, the author will draw a conclusion about Assessment Stoichiometry Answer Key. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Assessment Stoichiometry Answer Key.

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