

**AP[®] CHEMISTRY
2013 SCORING GUIDELINES**

Question 2

Answer the following questions involving the stoichiometry and thermodynamics of reactions containing aluminum species.



An electrolytic cell produces 235 g of $\text{Al}(l)$ according to the equation above.

(a) Calculate the number of moles of electrons that must be transferred in the cell to produce the 235 g of $\text{Al}(l)$.

$235 \text{ g Al} \times \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} = 8.71 \text{ mol Al}$ $\text{Al}^{3+} + 3 e^- \rightarrow \text{Al}, \text{ therefore, } 3 \text{ mol } e^- \text{ transferred per mol Al}$ $8.71 \text{ mol Al} \times \frac{3 \text{ mol } e^-}{1 \text{ mol Al}} = 26.1 \text{ mol } e^-$	<p>1 point is earned for the number of moles of Al.</p> <p>1 point is earned for correct stoichiometry and the number of moles of electrons.</p>
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(b) A steady current of 152 amp was used during the process. Determine the amount of time, in seconds, that was needed to produce the $\text{Al}(l)$.

$\text{charge} = \text{moles } e^- \times \text{Faraday's constant}$ $= 26.1 \text{ mol } e^- \times \frac{9.65 \times 10^4 \text{ C}}{1 \text{ mol } e^-} = 2.52 \times 10^6 \text{ C}$ $I = \frac{q}{t}$ $t = \frac{q}{I} = \frac{2.52 \times 10^6 \text{ C}}{152 \text{ C/s}} = 1.66 \times 10^4 \text{ s}$	<p>1 point is earned for the correct amount of charge transferred.</p> <p>1 point is earned for the correct time.</p>
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(c) Calculate the volume of $\text{CO}_2(g)$, measured at 301 K and 0.952 atm, that is produced in the process.

$\text{mol CO}_2 = 8.71 \text{ mol Al} \times \frac{3 \text{ mol CO}_2}{4 \text{ mol Al}} = 6.53 \text{ mol CO}_2$ $PV = nRT$ $V = \frac{nRT}{P} = \frac{(6.53 \text{ mol}) \left(0.0821 \frac{\text{L atm}}{\text{mol K}} \right) (301 \text{ K})}{0.952 \text{ atm}} = 1.70 \times 10^2 \text{ L CO}_2$	<p>1 point is earned for the number of moles of CO_2.</p> <p>1 point is earned for the volume of CO_2.</p>
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Roman Wölfel



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Table of Contents 2013 Ap Physics B Response Scoring Guidelines

1. Understanding the eBook 2013 Ap Physics B Response Scoring Guidelines
 - The Rise of Digital Reading 2013 Ap Physics B Response Scoring Guidelines
 - Advantages of eBooks Over Traditional Books
2. Identifying 2013 Ap Physics B Response Scoring Guidelines
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2013 Ap Physics B Response Scoring Guidelines
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2013 Ap Physics B Response Scoring Guidelines
 - Personalized Recommendations
 - 2013 Ap Physics B Response Scoring Guidelines User Reviews and Ratings
 - 2013 Ap Physics B Response Scoring Guidelines and Bestseller Lists
5. Accessing 2013 Ap Physics B Response Scoring Guidelines Free and Paid eBooks
 - 2013 Ap Physics B Response Scoring Guidelines Public Domain eBooks
 - 2013 Ap Physics B Response Scoring Guidelines eBook Subscription Services
 - 2013 Ap Physics B Response Scoring Guidelines Budget-Friendly Options

6. Navigating 2013 Ap Physics B Response Scoring Guidelines eBook Formats
 - ePub, PDF, MOBI, and More
 - 2013 Ap Physics B Response Scoring Guidelines Compatibility with Devices
 - 2013 Ap Physics B Response Scoring Guidelines Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2013 Ap Physics B Response Scoring Guidelines
 - Highlighting and Note-Taking 2013 Ap Physics B Response Scoring Guidelines
 - Interactive Elements 2013 Ap Physics B Response Scoring Guidelines
8. Staying Engaged with 2013 Ap Physics B Response Scoring Guidelines
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2013 Ap Physics B Response Scoring Guidelines
9. Balancing eBooks and Physical Books 2013 Ap Physics B Response Scoring Guidelines
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2013 Ap Physics B Response Scoring Guidelines
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2013 Ap Physics B Response Scoring Guidelines
 - Setting Reading Goals 2013 Ap Physics B Response Scoring Guidelines
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2013 Ap Physics B Response Scoring Guidelines
 - Fact-Checking eBook Content of 2013 Ap Physics B Response Scoring Guidelines
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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