

Begin your response to **QUESTION 1** on this page.**CHEMISTRY****SECTION II****Time—1 hour and 45 minutes****7 Questions**

Directions: Questions 1–3 are long free-response questions that require about 23 minutes each to answer and are worth 10 points each. Questions 4–7 are short free-response questions that require about 9 minutes each to answer and are worth 4 points each.

For each question, show your work for each part in the space provided after that part. Examples and equations may be included in your responses where appropriate. For calculations, clearly show the method used and the steps involved in arriving at your answers. You must show your work to receive credit for your answer. Pay attention to significant figures.

1. Answer the following questions related to manganese compounds.

(a) Manganese has several common oxidation states.

(i) Write the complete electron configuration for an Mn atom in the ground state.



(ii) When manganese forms cations, electrons are lost from which subshell first? Identify both the number and letter associated with the subshell.

4/s

A student performs an experiment to produce a manganese salt of unknown composition, $\text{Mn}_x\text{Cl}_y(\text{aq})$, and determine its empirical formula. The student places a sample of $\text{Mn}(s)$ in a beaker containing excess $\text{HCl}(\text{aq})$, as represented by the following equation.

**GO ON TO THE NEXT PAGE.**

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Princeton Review



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