



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Key Figures

- Figure 1: pH scale
- Figure 2: Ionization of water

Principles of Ionization

Acid dissociation is the process of ionization in which a substance dissociates into ions. This process is reversible.

Water dissociates into H_3O^+ and OH^- ions. The concentration of these ions is equal to the concentration of water molecules.

Strong acids dissociate completely. HCl dissociates into H^+ and Cl^- ions. H_2SO_4 dissociates into H^+ and SO_4^{2-} ions. HNO_3 dissociates into H^+ and NO_3^- ions.

Weak acids dissociate partially. CH_3COOH dissociates into H^+ and CH_3COO^- ions. H_2CO_3 dissociates into H^+ and HCO_3^- ions.

Strong bases dissociate completely. NaOH dissociates into Na^+ and OH^- ions.

Weak bases dissociate partially. NH_3 dissociates into NH_4^+ and OH^- ions.

Acid-Base Properties

<ul style="list-style-type: none">strong acidstrong baseweak acidweak baseacidbaseconjugate acidconjugate base	<ul style="list-style-type: none">strong acidstrong baseweak acidweak baseacidbaseconjugate acidconjugate base	<ul style="list-style-type: none">strong acidstrong baseweak acidweak baseacidbaseconjugate acidconjugate base
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