

Multiple Choice

Select the alternative that best completes the statement or answers the question. What is the most common chemical gas used?

1. How many moles of hydrogen gas are in a cylinder at 100 atm and 25°C?
- ☐ 1
☒ 2
☐ 3
☐ 4

2. At 25°C, the K_{sp} of AgCl is 1.8×10^{-10} . What is the molar solubility of AgCl in water?
- ☒ $1.3 \times 10^{-5} \text{ M}$
☐ $1.3 \times 10^{-10} \text{ M}$
☐ $1.3 \times 10^{-5} \text{ M}$
☐ $1.3 \times 10^{-10} \text{ M}$

3. The standard enthalpy change for the reaction is:
- ☐ $\Delta H^\circ_{\text{rxn}} = 100 \text{ kJ/mol}$
☐ $\Delta H^\circ_{\text{rxn}} = -100 \text{ kJ/mol}$
☒ $\Delta H^\circ_{\text{rxn}} = 200 \text{ kJ/mol}$
☒ $\Delta H^\circ_{\text{rxn}} = -200 \text{ kJ/mol}$

4. For the reaction $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$, $\Delta H^\circ_{\text{rxn}} = -184.6 \text{ kJ/mol}$. How much heat is released when 1 mole of H_2 reacts with 1 mole of Cl_2 ?
- ☐ 184.6 kJ
☐ 92.3 kJ
☐ 369.2 kJ
☒ 184.6 kJ

5. What is the pH of a 0.1 M solution of HCl ?
- ☐ 1.0
☒ 1.0
☐ 1.0
☐ 1.0

6. What is the molar solubility of AgCl in water?
- ☐ $1.3 \times 10^{-5} \text{ M}$
☐ $1.3 \times 10^{-10} \text{ M}$
☐ $1.3 \times 10^{-5} \text{ M}$
☒ $1.3 \times 10^{-10} \text{ M}$

7. A solution of HCl is prepared by adding 100 mL of 1.0 M HCl to 100 mL of 1.0 M HCl . What is the pH of the resulting solution?
- ☐ 1.0
☒ 1.0
☐ 1.0
☐ 1.0

8. For the reaction $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$, $\Delta H^\circ_{\text{rxn}} = -184.6 \text{ kJ/mol}$. How much heat is released when 1 mole of H_2 reacts with 1 mole of Cl_2 ?
- ☒ 184.6 kJ
☐ 92.3 kJ
☐ 369.2 kJ
☐ 184.6 kJ

9. At 25°C, the K_{sp} of AgCl is 1.8×10^{-10} . What is the molar solubility of AgCl in water?
- ☐ $1.3 \times 10^{-5} \text{ M}$
☐ $1.3 \times 10^{-10} \text{ M}$
☐ $1.3 \times 10^{-5} \text{ M}$
☒ $1.3 \times 10^{-10} \text{ M}$

10. What is the pH of a 0.1 M solution of HCl ?
- ☐ 1.0
☒ 1.0
☐ 1.0
☐ 1.0

11. How much heat is released when 1 mole of H_2 reacts with 1 mole of Cl_2 ?

Multiple Choice Stoichiometry Test

John T. Moore, Richard H. Langley



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