

Module #3: Atoms and Molecules

Pre-Lab Prep: Key Definitions and Math Necessary for Labs

The Law of Mass Conservation: Matter cannot be created or destroyed; it can only change forms.

Be sure you understand applications of this law such as OYO 3.1: A chemist burns 15.4 kg of wood in a fireplace. The chemist collects all of the substances produced by the fire. He measures the mass of ash produced to be 925 grams, the mass of carbon dioxide produced to be 15.14 gk and the mass of water to be 2.12 kg. How much oxygen was used by the wood while it burned?

Element: Any substance that cannot be decomposed into less massive substances.

Be sure to have a Periodic Table of Elements available for quick, easy, frequent reference.

Be able understand the information in the Periodic chart:

- Abbreviation of the element's Latin or English name
- Have memorized the first 20 elements on the chart (H through Ca)
- How to identify metals vs. nonmetals on the chart, and the one exception to this rule (H: hydrogen)

A compound is: A substance that can be decomposed into elements by chemical means.

Law of Definite Proportions – The proportion of elements in any compound is always the same. Be sure you understand Example 3.2 on page 80.

Law of Multiple Proportions: If two elements combine to form different compounds, the ratio of masses of the *second* element that react with the fixed mass of the *first* element will be a simple, whole-number ratio.

Dalton's Atomic Theory: a) All elements are composed of small, invisible particles called "atoms"; b) All atoms of the same element have exactly the same properties; c) Atoms of different elements have different properties; and d) Compounds are formed when atoms are joined together. Since atoms are indivisible, they can only join together in simple, whole-number ratios.

What is the difference between a mixture of atoms and a compound? A mixture has free-flowing atoms, but a compound has molecules with the atoms bonded together.

What is a chemical symbol and a chemical formula? Give an example with water. H / H₂O – the subscript number tells how many atoms of the preceding symbol are in the molecule.

What is the difference between ionic and covalent matter? It is ionic if, when dissolved in water, it conducts electricity. Compounds made up of at least one metal atom and at least one nonmetal atom is ionic. All other matter is covalent, and covalent compounds are made up of solely nonmetal atoms.

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