

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the quantity indicated that are indicated in the problem. Be sure to include units whenever a unit is indicated or implied.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of those species present in water that equal to _____.
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19 2 Hydrogen Ions And Acidity Answer Key

Karin Nielsen-Saines



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version Units I and II Chapters 1-16 **Biology** Teresa Audesirk, Gerald Audesirk, Bruce E. Byers, 2004-06 For students without an Internet connection all questions and review materials from the Companion Website are included in the printed Student Study Companion **Applied Science & Technology Index**, 1983

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 - Chapter 2: Essential Elements of 19 2 Hydrogen Ions And Acidity Answer Key
 - Chapter 3: 19 2 Hydrogen Ions And Acidity Answer Key in Everyday Life
 - Chapter 4: 19 2 Hydrogen Ions And Acidity Answer Key in Specific Contexts
 - Chapter 5: Conclusion
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