

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)
- a) Translate the problem to chemical equation.



Reactant side
4 Fe atoms
4 S atoms
 $7(2) = 14$ O atoms

Product side
4 Fe atoms
4 S atoms
 $2(3) + 4(2) = 14$ O atoms

- b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

$m \text{ FeS} = 2.0 \text{ g}$
Req'd. = $m \text{ O}_2$

Solution:

molar mass of $\text{FeS} = 88 \text{ g/mol}$
mole ratio = 4 mol FeS : 7 mol O_2
molar mass of $\text{O}_2 = 32 \text{ g/mol}$

$$m_{\text{O}_2} = 2.0 \text{ g FeS} \left(\frac{1 \text{ mol FeS}}{88 \text{ g FeS}} \right) \left(\frac{7 \text{ mol O}_2}{4 \text{ mol FeS}} \right) \left(\frac{32 \text{ g O}_2}{1 \text{ mol O}_2} \right) = 1.3 \text{ g O}_2$$

2. Rusting involves the reaction of iron, Fe with oxygen, O_2 to form iron (III) oxide, Fe_2O_3 . (20pts)

- a) Write the balanced chemical equation.



Reactant side
4 Fe atoms
 $3(2) = 6$ O atoms

Product side
4 Fe atoms
 $2(3) = 6$ O atoms

- b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

$$4 \text{ mol Fe} : 3 \text{ mol O}_2$$

Product to each reactant:

$$2 \text{ mol Fe}_2\text{O}_3 : 4 \text{ mol Fe}$$

Answers To Stoichiometry With 3 Problems

SA Dillow

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