

Redox Reactions - Class 11 Chemistry Worksheet (With Answers)

Section A: Find the Oxidation Number - Answers

1. H_2O_2 : Oxidation number of O is -1 (peroxide).
2. Mn in KMnO_4 : +7
3. Cr in $\text{Cr}_2\text{O}_7^{2-}$: +6
4. S in H_2SO_4 : +6
5. Cl in ClO_3^- : +5
6. Fe in Fe_3O_4 : Average +8/3 (as Fe^{2+} and Fe^{3+} both are present).
7. N in NH_4^+ : -3
8. C in HCOOH : +2 (H-C-O is +2 and -OH contributes no change to that C).
9. Cu in $\text{Cu}(\text{NO}_3)_2$: +2
10. O in OF_2 : +2 (because F is more electronegative and takes -1).

Section B: Identify Oxidizing and Reducing Agents - Answers

11. Zn is reducing agent, H^+ (from H_2SO_4) is oxidizing agent.
12. Cu is reducing agent, Ag^+ is oxidizing agent.
13. Cl_2 is oxidizing agent, Br^- is reducing agent.
14. CO is reducing agent, Fe_2O_3 is oxidizing agent.
15. MnO_2 is oxidizing agent, HCl (Cl^-) is reducing agent.

Section C: Balanced Redox Reactions (Acidic Medium) - Answers

16. $\text{Cr}_2\text{O}_7^{2-} + 6\text{Fe}^{2+} + 14\text{H}^+ \rightarrow 2\text{Cr}^{3+} + 6\text{Fe}^{3+} + 7\text{H}_2\text{O}$
17. $2\text{MnO}_4^- + 5\text{SO}_3^{2-} + 6\text{H}^+ \rightarrow 2\text{Mn}^{2+} + 5\text{SO}_4^{2-} + 3\text{H}_2\text{O}$
18. $2\text{NO}_3^- + 10\text{I}^- + 12\text{H}^+ \rightarrow 2\text{NO} + 5\text{I}_2 + 6\text{H}_2\text{O}$
19. $\text{H}_2\text{O}_2 + 2\text{Fe}^{2+} + 2\text{H}^+ \rightarrow 2\text{Fe}^{3+} + 2\text{H}_2\text{O}$

Answers To Redox Webquest

Malika Ammam



Answers To Redox Webquest:

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Malika Ammam, 2017-11-25 This short book part 1 summarizes some basics of redox chemistry electrochemistry in solution Each section of the book contains concise and clear referenced summaries of redox concepts Over 250 answered questions and problems are provided to further clarify the discussed principles

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Table of Contents Answers To Redox Webquest

1. Understanding the eBook Answers To Redox Webquest
 - The Rise of Digital Reading Answers To Redox Webquest
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers To Redox Webquest
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers To Redox Webquest
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers To Redox Webquest
 - Personalized Recommendations
 - Answers To Redox Webquest User Reviews and Ratings
 - Answers To Redox Webquest and Bestseller Lists
5. Accessing Answers To Redox Webquest Free and Paid eBooks
 - Answers To Redox Webquest Public Domain eBooks
 - Answers To Redox Webquest eBook Subscription Services
 - Answers To Redox Webquest Budget-Friendly Options
6. Navigating Answers To Redox Webquest eBook Formats

- ePub, PDF, MOBI, and More
- Answers To Redox Webquest Compatibility with Devices
- Answers To Redox Webquest Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers To Redox Webquest
 - Highlighting and Note-Taking Answers To Redox Webquest
 - Interactive Elements Answers To Redox Webquest
- 8. Staying Engaged with Answers To Redox Webquest
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers To Redox Webquest
- 9. Balancing eBooks and Physical Books Answers To Redox Webquest
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers To Redox Webquest
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers To Redox Webquest
 - Setting Reading Goals Answers To Redox Webquest
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers To Redox Webquest
 - Fact-Checking eBook Content of Answers To Redox Webquest
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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