

## Reviewing Content

16.1. The solvent is the substance in which the solute is dissolved.

16.2. Dissolved solutes of the solvent interact with the solute particles, providing enough force to overcome gravity.

16.3. Osmolarity is the measure of a substance that dissolves in a given quantity of solvent at a given amount of solute and a constant temperature and pressure. **nonideal solutio**n: a solution that contains less solute than a saturated solution at a given temperature and pressure. **oversaturated solutio**n: a solution that contains more solute than a saturated solution at a given temperature and pressure. **unsaturated solutio**n: a solution that contains less solute than a saturated solution at a given temperature and pressure. **saturation**: the point at which no more solute can dissolve in a given amount of solvent.

16.4. Particles of solute crystallize.

16.5. If there were unsaturated solute, the excess solute would enter into a supersaturated solution.

16.6.  $0.50 \times 10^{-2} \text{ g AgNO}_3$ .

16.7. Osmolarity increases with pressure.

16.8. a.  $1.6 \times 10^{-2} \text{ g/L}$   
b.  $1.7 \times 10^{-2} \text{ g/L}$

16.9. Osmolarity and concentration are relative terms and are not quantitative. Osmolarity provides the exact number of moles of solute per liter of solution.

16.10. Osmolarity is the measure of moles of solute dissolved in one liter of solution.  
a.  $1.00 \text{ M NaCl}$   
b.  $0.20 \times 10^{-2} \text{ M MgCl}_2$

16.11.  $2.000 \times 10^{-2} \text{ mol}$ .

16.12. a.  $1.00 \times 10^{-2} \text{ mol NaCl}$ ,  $2.00 \text{ g NaCl}$   
b.  $1.00 \text{ mol KNO}_3$ ,  $1.00 \times 10^{-2} \text{ g KNO}_3$   
c.  $2.00 \times 10^{-2} \text{ mol CaCl}_2$ ,  $2.00 \text{ g CaCl}_2$

16.13. a.  $2.00 \times 10^{-2} \text{ g NaCl}$   
b.  $2.00 \text{ g MgCl}_2$

16.14. a. 100% EtOH ethanol  
b. 90.0% EtOH isopropyl alcohol

16.15. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because solutes of molten form surround solute particles, reducing the amount of solute molecules that have sufficient energy to escape the solution. Relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent molecules, so more kinetic energy must be withdrawn from a solution than is necessary. This lowers the freezing point of the solution.

16.16. a. 100% EtOH  
b.  $1.00 \text{ M NaCl}$   
c.  $0.1000 \text{ M MgCl}_2$

16.17. The effective osmolarity of the  $\text{Ca}(\text{NO}_3)_2$  solution is 2.0. The effective osmolarity of the  $\text{NaNO}_3$  solution is 2.0.

16.18. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.

16.19. The salt lowers the freezing point of the ice-water cooling mixture.

16.20. 1.0M solutions: 1 mol of solute in 1 L of solution; 1 mol solute: 1 mol of solute in 0.998 L of solution.

16.21.  $16.00 \times 10^{-2} \text{ g H}_2\text{O}$  or  $32.00 \text{ g C}_2\text{H}_5\text{OH}$ .

16.22. a.  $-1.00^\circ\text{C}$   
b.  $-0.54^\circ\text{C}$

16.23. a.  $-0.1^\circ\text{C}$   
b.  $-0.01^\circ\text{C}$   
c.  $-0.07^\circ\text{C}$

## Understanding Concepts

16.24. a. The boiling-point depression is not as great for solution B; solute B causes fewer like particles as many particles in solution A.

# Answers To Prentice Hall Chemistry

**University Interscholastic League  
(Tex.)**

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