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## Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \rightarrow \text{LiCl}$ $\text{Li} \rightarrow \text{Li}^+ \text{:Cl}^-$	Li	1	1 transfer	$\text{Li}^+$
	Cl	7	1 gain	$\text{Cl}^-$
2) $\text{Ca} + \text{O} \rightarrow \text{CaO}$ $\text{Ca} \cdot \text{:O} \rightarrow \text{Ca}^{+2} \cdot \text{:O}^{-2}$	Ca	2	2 transfer	$\text{Ca}^{+2}$
	O	6	2 gain	$\text{O}^{-2}$
3) $\text{Be} + \text{F} \rightarrow \text{BeF}_2$ $\text{Be} \cdot + \text{:F}^- \rightarrow \text{Be}^{+2} \cdot \text{:F}^{-2}$	Be	2	2 transfer	$\text{Be}^{+2}$
	F	7	each F gains 1 e <sup>-</sup>	$\text{F}^-$
4) $\text{Mg} + \text{S} \rightarrow \text{MgS}$ $\text{Mg} \cdot + \text{:S}^- \rightarrow \text{Mg}^{+2} \cdot \text{:S}^{-2}$	Mg	2	2 transfer	$\text{Mg}^{+2}$
	S	6	2 gain	$\text{S}^{-2}$
5) $\text{K} + \text{F} \rightarrow \text{KF}$ $\text{K} \cdot \text{:F}^- \rightarrow \text{K}^+ \cdot \text{:F}^{-1}$	K	1	1 transfer	$\text{K}^+$
	F	7	1 gain	$\text{F}^{-1}$

# Answers To Ionic Bonds

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