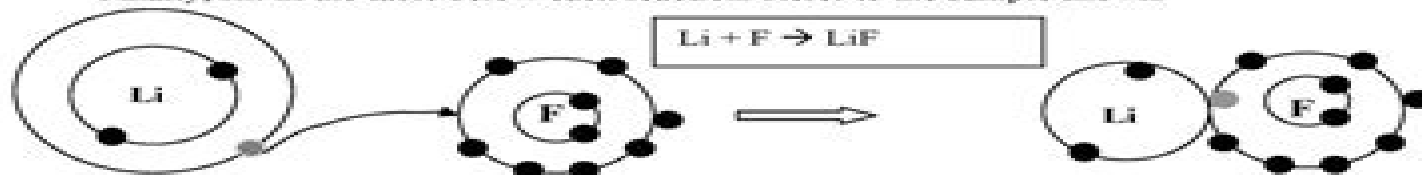


Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms Valence electrons Electron transfer from/to each atom Ions formed in the product

Li
F

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$ $\text{Li} \cdot + \cdot \ddot{\text{Cl}} : \rightarrow \text{Li}^+ : \ddot{\text{Cl}} : ^-$	Li	1	1 transfer	Li^+
	Cl	7	1 gain	Cl^-
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$ $\text{Ca} \cdot + \cdot \ddot{\text{O}} : \rightarrow \text{Ca}^{2+} : \ddot{\text{O}} : ^{2-}$	Ca	2	2 transfer	Ca^{2+}
	O	6	2 gain	O^{2-}
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$ $\text{Be} \cdot + \cdot \ddot{\text{F}} : + \cdot \ddot{\text{F}} : \rightarrow \text{Be}^{2+} : \ddot{\text{F}} : ^- : \ddot{\text{F}} : ^-$	Be	2	2 transfer	Be^{2+}
	F	7	each F gains 1 e	F^-
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$ $\text{Mg} \cdot + \cdot \ddot{\text{S}} : \rightarrow \text{Mg}^{2+} : \ddot{\text{S}} : ^{2-}$	Mg	2	2 transfer	Mg^{2+}
	S	6	2 gain	S^{2-}
5) $\text{K} + \text{F} \Rightarrow \text{KF}$ $\text{K} \cdot + \cdot \ddot{\text{F}} : \rightarrow \text{K}^+ : \ddot{\text{F}} : ^-$	K	1	1 transfer	K^+
	F	7	1 gain	F^-

Answers To Ionic Bonds

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