

## Chemthink: CHEMICAL REACTIONS

- a) Go to [www.chemthink.com](http://www.chemthink.com). Log in using your user name and password. (If you forgot, you need to WAIT QUIETLY until your teacher is done introducing the lesson.)
- b) Scroll to **CHEMICAL REACTIONS** then the tutorial under the Chemical Reactions menu.
- c) Complete the **tutorial**. As you complete the tutorial, write down the answers to the questions below.
- d) Once you have successfully completed the tutorial and this worksheet, complete the **Question Set**.

### CHEMICAL REACTIONS TUTORIAL QUESTIONS:

- Starting materials in a chemical reaction are called \_\_\_\_\_.
- The ending materials in a chemical reaction are called \_\_\_\_\_.
- The arrow indicates a \_\_\_\_\_ has taken place.
- All reactions have one thing in common: there is a \_\_\_\_\_ of chemical bonds.
- Chemical reactions always involve \_\_\_\_\_, old bonds, \_\_\_\_\_, new bonds, or both.
- In all reactions we still have all of the \_\_\_\_\_ at the end that we had at the start.
- In every reaction there can never be any \_\_\_\_\_ atoms or \_\_\_\_\_ in the atoms that are already there.
- Chemical reactions only \_\_\_\_\_ in the atoms that are already there.
- Let's represent a reaction on paper. For example, hydrogen gas ( $H_2$ ) reacts with oxygen gas ( $O_2$ ) to form water ( $H_2O$ ):



If we use only the atoms shown, we'd have \_\_\_\_\_ atoms of H and \_\_\_\_\_ atoms of O as reactants. This would make \_\_\_\_\_ molecule of  $H_2O$ , but we'd have \_\_\_\_\_ atom of O leftover. However, this reaction only makes  $H_2O$ .

Remember: reactions are not limited to 1 molecule each of reactants. We can use as many as we need to balance the chemical equation.

#### A balanced chemical reaction shows:

- a) What atoms are present before (in the reactants) and after (in the products)
  - b) **How many** of each reactant and product is present before and after.
10. So to make  $H_2O$  from oxygen gas and hydrogen gas, the balanced equation would be:



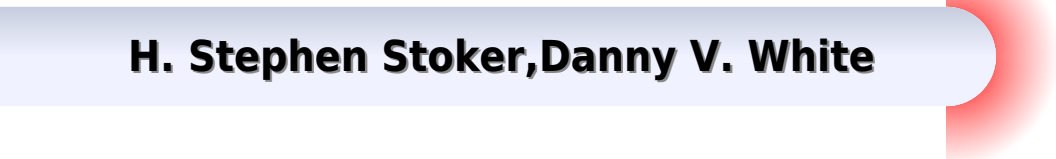
Which is the same as:

| # of atoms in Reactants | Element | # of atoms in Products |
|-------------------------|---------|------------------------|
|                         | H       |                        |
|                         | O       |                        |

- This idea is called the \_\_\_\_\_.
- There must be the same \_\_\_\_\_ and the same number of \_\_\_\_\_ before the reaction (in the reactants) and after the reaction (in the products).
- What is the balanced equation for this reaction? \_\_\_\_\_  $Cu$  + \_\_\_\_\_  $O_2 \rightarrow$  \_\_\_\_\_  $CuO$

# Answers To Chemthink Questions Tutorial Test

**H. Stephen Stoker, Danny V. White**



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