

Molarity

(How Concentrated Is It?)

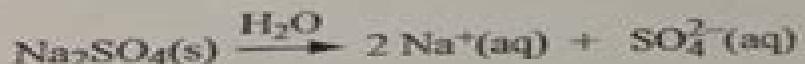
Water is the most common solvent, and we will focus on aqueous solutions. However, the terms applied herein also apply when other solvents are used.

Solutes are categorized according to their ability to affect the electrical conductivity of the solution upon dissolution of the solute in water.

- An aqueous solution of a **strong electrolyte** conducts electricity well.
- An aqueous solution of a **nonelectrolyte** does not conduct electricity.
- An aqueous solution of a **weak electrolyte** conducts electricity poorly.
- Water from most common sources (tap water, rain water, water in lakes and rivers, etc.) always contains strong electrolytes and conducts electricity well.

Model 1: Strong Electrolytes.

Many ionic compounds are strong electrolytes. When an ionic compound dissolves in water the cations and anions are separated. In the solution, the cations and anions are surrounded by water molecules. Sodium sulfate and sodium chloride are both strong electrolytes.



The "(aq)" after the ion indicates that each ion of that type is surrounded by several water molecules. It has been discovered that it is the mobile, charged particles in the solution that carry the electric current.

Answers To Chemactivity 32 Molarity

Richard S. Moog, John J. Farrell

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