

Exercises



- Provide an expression for K_p for reaction 1. Why does a term for C(graphite) not appear in the K_p expression for the reaction of Table 1?
- Calculate the equilibrium constant for the reaction using tabulated data.

$$\Delta_r H^\circ = -48.4 \text{ kJ/mol}$$



$$\Delta_r S^\circ = -87 \text{ J/K/mol}$$

Write the equilibrium constant for this reaction in terms of the partial pressures.

- Calculate the equilibrium constant for

$$\Delta_r G_f^\circ (\text{CaO}) = -646.3 \text{ kJ/mol}$$



$$\Delta_r G_f^\circ (\text{CO}_2) = -394.36 \text{ kJ/mol}$$

from tabulated data.

$$\Delta_r G_f^\circ (\text{CaO}) = -128.77 \text{ kJ/mol}$$

- The reaction for the formation of nitrosyl chloride



was studied at 25°C. The pressures at equilibrium were found to be

$$P_{\text{NOCl}} = 1.2 \text{ atm}$$

$$P_{\text{NO}} = 5.0 \times 10^{-2} \text{ atm}$$

$$P_{\text{Cl}_2} = 3.0 \times 10^{-1} \text{ atm}$$

- Calculate the value of K_p for this reaction at 25°C.
- If 1.0 atm of NOCl(g) was placed into an evacuated container would a reaction occur?
- If 1.0 atm each of NOCl(g), Cl₂(g) and NO(g) were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?
- If 1.0 atm of NOCl(g), 5.0×10^{-2} atm of NO and 3.0×10^{-1} atm of Cl₂ were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?

Show all calculations and explain each answer.

Answers To Chemactivity 4

Richard Samuel Moog

Answers To Chemactivity 4:

Chemistry Richard S. Moog, John J. Farrell, 2014-01-13 Chemistry A Guided Approach 6th Edition follows the underlying principles developed by years of research on how readers learn and draws on testing by those using the POGIL methodology. This text follows inquiry based learning and correspondingly emphasizes the underlying concepts and the reasoning behind the concepts. This text offers an approach that follows modern cognitive learning principles by having readers learn how to create knowledge based on experimental data and how to test that knowledge.

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Organic Chemistry, a Guided Inquiry Andrei Straumanis, 2004 Process Oriented Guided Inquiry Learning (POGIL) Richard Samuel Moog, 2008 POGIL is a student centered group learning pedagogy based on current learning theory. This volume describes POGIL's theoretical basis, its implementations in diverse environments, and evaluation of student outcomes.

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 Prentice Hall Chemistry, 2000

Chemists' Guide to Effective Teaching Norbert J. Pienta, Melanie M. Cooper, Thomas J. Greenbowe, 2005 For courses in Methods of Teaching Chemistry. Useful for new professors, chemical educators, or students learning to teach chemistry. Intended for anyone who teaches chemistry or is learning to teach it, this book examines applications of learning theories, presenting actual techniques and practices that respected professors have used to implement and achieve their goals. Each chapter is written by a chemist who has expertise in the area and who has experience in applying those ideas in their classrooms. This book is a part of the Prentice Hall Series in Educational Innovation for Chemistry.

Answers to Revision Questions for National 5 Chemistry Douglas Buchanan, James Melrose, A. R. Ross, 2013

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- Chapter 2: Essential Elements of Answers To Chemactivity 4
- Chapter 3: Answers To Chemactivity 4 in Everyday Life
- Chapter 4: Answers To Chemactivity 4 in Specific Contexts
- Chapter 5: Conclusion

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