

## Exercises



1. Provide an expression for  $K_P$  for reaction 1. Why does a term for C(graphite) not appear in the  $K_P$  expression for the reaction of Table 1?
2. Calculate the equilibrium constant for the reaction using tabulated data.

$$\Delta_r H^\circ = -484 \text{ kJ/mol}$$

$$\Delta_r S^\circ = -89 \text{ kJ/mol}$$



Write the equilibrium constant for this reaction in terms of the partial pressures.

3. Calculate the equilibrium constant for

$$\Delta_r G_f^\circ(\text{CaO}) = -604.3 \text{ kJ/mol}$$

$$\Delta_r G_f^\circ(\text{CO}_2) = -394.36 \text{ kJ/mol}$$

from tabulated data.

$$\Delta_r G_f^\circ(\text{CaCO}_3) = -1128.77 \text{ kJ/mol}$$



9. The reaction for the formation of nitrosyl chloride



was studied at 25°C. The pressures at equilibrium were found to be

$$P_{\text{NOCl}} = 1.2 \text{ atm}$$

$$P_{\text{NO}} = 5.0 \times 10^{-2} \text{ atm}$$

$$P_{\text{Cl}_2} = 3.0 \times 10^{-1} \text{ atm}$$

- a) Calculate the value of  $K_P$  for this reaction at 25°C.
- b) If 1.0 atm of  $\text{NOCl}(\text{g})$  was placed into an evacuated container would a reaction occur?
- c) If 1.0 atm each of  $\text{NOCl}(\text{g})$ ,  $\text{Cl}_2(\text{g})$  and  $\text{NO}(\text{g})$  were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?
- d) If 1.0 atm of  $\text{NOCl}(\text{g})$ ,  $5.0 \times 10^{-3}$  atm of  $\text{NO}$  and  $3.0 \times 10^{-1}$  atm of  $\text{Cl}_2$  were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?

Show all calculations and explain each answer.

# Answers To Chemactivity 4

**Richard Samuel Moog**



#### Answers To Chemactivity 4:

*Chemistry* Richard S. Moog, John J. Farrell, 2014-01-13 Chemistry A Guided Approach 6th Edition follows the underlying principles developed by years of research on how readers learn and draws on testing by those using the POGIL methodology This text follows inquiry based learning and correspondingly emphasizes the underlying concepts and the reasoning behind the concepts This text offers an approach that follows modern cognitive learning principles by having readers learn how to create knowledge based on experimental data and how to test that knowledge     Introductory Chemistry Michael P.

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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998     **Prentice Hall Chemistry**, 2000

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The book delves into Answers To Chemactivity 4. Answers To Chemactivity 4 is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Answers To Chemactivity 4, encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to Answers To Chemactivity 4
- Chapter 2: Essential Elements of Answers To Chemactivity 4
- Chapter 3: Answers To Chemactivity 4 in Everyday Life
- Chapter 4: Answers To Chemactivity 4 in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Answers To Chemactivity 4. The first chapter will explore what Answers To Chemactivity 4 is, why Answers To Chemactivity 4 is vital, and how to effectively learn about Answers To Chemactivity 4.
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6. In chapter 5, this book will draw a conclusion about Answers To Chemactivity 4. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Answers To Chemactivity 4.

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