

## 7-1 Review and Reinforcement

### Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative	noble gas	octet	electron configuration
positive	electrons	ions	polyatomic
crisscross	anion	cation	empirical formula

1. An ionic bond is an attraction between oppositely charged \_\_\_\_\_.
2. Anions have a \_\_\_\_\_ charge.
3. An atom becomes an ion by losing or gaining \_\_\_\_\_.
4. The \_\_\_\_\_ rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) \_\_\_\_\_.
6. The atoms that make up \_\_\_\_\_ ions are bonded together by covalent bonds.
7. The \_\_\_\_\_ method can be used to write the formula for an ionic compound.
8. The \_\_\_\_\_ of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

9. Calcium becomes a monatomic cation by gaining two electrons.  
\_\_\_\_\_
10. A cation has a positive charge.  
\_\_\_\_\_
11. A binary ionic compound contains only one kind of cation and one kind of anion.  
\_\_\_\_\_
12. Monatomic ions consist of more than one atom.  
\_\_\_\_\_
13. The Lewis dot diagram for chlorine, a group 7A element, has six electrons.  
\_\_\_\_\_
14. A great deal of energy is produced when an ionic compound is formed.  
\_\_\_\_\_

# 7 Review And Reinforcement Ionic Bonding

**S Ben Porath**



## 7 Review And Reinforcement Ionic Bonding:

*NBS Special Publication* ,1978      *Catalog of National Bureau of Standards Publications, 1966-1976: Key word index*  
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