

7.1

IONS

Section Review

Objectives

- 1. Understand the structure of ionic compounds using periodic trends
- 2. Importance of crystal
- 3. Molecular formula mass
- 4. Empirical formula mass

Vocabulary

- 1. ionic compound
- 2. crystal lattice structure
- 3. formula mass
- 4. crystal mass

Part A Completion

Use the information on ions to determine the formula of the compound. Write the formula of the compound in the space provided.

1. A metal with a valence of +2 and a nonmetal with a valence of -3 combine to form a compound. The formula of the compound is M_2X_3 . The metal is Ca and the nonmetal is N . The formula of the compound is Ca_3N_2 . The metal is Ca and the nonmetal is N . The formula of the compound is Ca_3N_2 .

2. A metal with a valence of +2 and a nonmetal with a valence of -3 combine to form a compound. The formula of the compound is M_2X_3 . The metal is Ca and the nonmetal is N . The formula of the compound is Ca_3N_2 . The metal is Ca and the nonmetal is N . The formula of the compound is Ca_3N_2 .

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- 6. Ca_3N_2
- 7. Ca_3N_2
- 8. Ca_3N_2
- 9. Ca_3N_2
- 10. Ca_3N_2
- 11. Ca_3N_2
- 12. Ca_3N_2
- 13. Ca_3N_2
- 14. Ca_3N_2
- 15. Ca_3N_2
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Michal Rosen-Zvi



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