

## 9-2 Review and Reinforcement

### Chemical Equations

Refer to the unbalanced chemical equation below to answer each of the following questions. Write your answers in the space provided.



1. Write the names of the reactants in this equation.

Aluminum nitrate + Iron (II) chloride

2. Write the names of the products in this equation.

Iron (II) nitrate + aluminum chloride

3. What would the coefficient of  $\text{Al(NO}_3\text{)}_3$  be in the balanced equation? Two

4. What would the coefficient of  $\text{Fe(NO}_3\text{)}_2$  be in the balanced equation? Three

5. What would the coefficient of  $\text{AlCl}_3$  be in the balanced equation? Two

6. How many oxygen atoms would there be on each side of the balanced equation? 18

Answer each of the following questions in the space provided.

7. What should you do if an equation you have written cannot be balanced?

Make sure reactants + products are written correctly, try again, ask for help

8. How is balancing a chemical reaction related to the law of conservation of matter?

mass of reactants must = mass of products

9. Describe the correct way to balance an equation.

add only coefficients to keep charge-balance  
or subscripts

10. How does a chemical equation relate the macroscopic, observable reaction to the interactions occurring at the microscopic level?

Subatomic

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**K Payea**

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## Table of Contents 2 Review And Reinforcement Answers

1. Understanding the eBook 2 Review And Reinforcement Answers
  - The Rise of Digital Reading 2 Review And Reinforcement Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying 2 Review And Reinforcement Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 2 Review And Reinforcement Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2 Review And Reinforcement Answers
  - Personalized Recommendations
  - 2 Review And Reinforcement Answers User Reviews and Ratings
  - 2 Review And Reinforcement Answers and Bestseller Lists
5. Accessing 2 Review And Reinforcement Answers Free and Paid eBooks
  - 2 Review And Reinforcement Answers Public Domain eBooks
  - 2 Review And Reinforcement Answers eBook Subscription Services
  - 2 Review And Reinforcement Answers Budget-Friendly Options
6. Navigating 2 Review And Reinforcement Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - 2 Review And Reinforcement Answers Compatibility with Devices
  - 2 Review And Reinforcement Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 2 Review And Reinforcement Answers
  - Highlighting and Note-Taking 2 Review And Reinforcement Answers
  - Interactive Elements 2 Review And Reinforcement Answers
8. Staying Engaged with 2 Review And Reinforcement Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 2 Review And Reinforcement Answers

9. Balancing eBooks and Physical Books 2 Review And Reinforcement Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 2 Review And Reinforcement Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 2 Review And Reinforcement Answers

- Setting Reading Goals 2 Review And Reinforcement Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 2 Review And Reinforcement Answers

- Fact-Checking eBook Content of 2 Review And Reinforcement Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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