

## Chapter 7 Chemical Reactions

**Section 7.1 Describing Reactions****(pages 192–198)**

*This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry.*

**Reading Strategy (page 192)**

**Monitoring Your Understanding:** Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn. After reading, state what you learned about each item you listed. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Expect to Learn	What I Learned
Answers may vary. Possible answers: How to balance chemical equations	Answers may vary. Possible answers: An unbalanced equation can be balanced by changing the coefficients.
How to convert from moles to moles	The mass of a substance can be converted to moles by using the molar mass as a conversion factor.

**Chemical Equations (pages 192–193)**

- Is the following sentence true or false? The new substances formed as a result of a chemical reaction are called products. True
- Circle the letter of each sentence that is a correct interpretation of the chemical equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ .
  - Carbon and oxygen react and form carbon monoxide.
  - ☒ Carbon and oxygen react and form carbon dioxide.
  - Carbon dioxide yields carbon and oxygen.
  - ☒ The reaction of carbon and oxygen yields carbon dioxide.
- Is the following sentence true or false? The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction. True
- Circle the letter of the correct answer. According to the equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ , how many carbon atoms react with 14 molecules of oxygen to form 14 molecules of carbon dioxide?
  - 1
  - 7
  - ☒ 14
  - 28
- In the reaction represented by the equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ , the mass of carbon dioxide produced equals the total mass of carbon and oxygen that reacted.

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**J Dewey**



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**Focus on Physical Science California Edition** Michael J. Padilla, 2007      **Prentice Hall Science Explorer** Michael J. Padilla, 2002      *Theoretical and Computational Research in the 21st Century* Nazmul Islam, 2014-10-28 This book focuses mainly on the recent developments of all types of theoretical mathematical and computational conceptions as well as modelling and simulation of specific research themes covering all scientific and technical disciplines from chemistry physics and engineering to biology and medicine The book contains timely reviews and research covering fundamental and applied research aspects in all disciplines of natural sciences including their historical representations and philosophical perspectives The book discusses the fact that the largest and smallest values of the Fukui function and local softness do not necessarily correspond to the softness and hardness regions of the molecules such as porphyrins The authors have adopted two popular calculation procedures for this venture One is the very old Hückel molecular orbital calculation and the other is one of the best semi empirical AM1 procedures for such systems Our finding is that neither the Fukui functions nor the local softnesses can predict the preferred donor sites of porphyrins toward metal ions      **Pearson's Chemical Analysis of Foods** Harold Egan, Ronald S. Kirk, Ronald Sawyer, 1981      Applied Kinetics and Chemical Reaction Engineering Robert L. Goring, Vern W. Weekman, 1967      **The Pharmaceutical Journal and Pharmacist**, 1918      **Annual Report** New Jersey. Board of Agriculture, 1889      **Annual Report of the State Board of Agriculture** New Jersey. State Board of Agriculture, 1888      **Reports** New Jersey. State Board of Agriculture, 1889      **The Druggists' Circular and Chemical Gazette**, 1878 Includes Red book price list section title varies slightly issued semiannually 1897 1906      *Pharmaceutical Journal*; , 1918      The Burning Answer Keith Barnham, 2014-05-15 Our civilisation stands on the brink of catastrophe Our thirst for energy has led to threats from global warming nuclear disaster and conflict in oil rich countries We are running out of options Solar power Keith Barnham argues is the answer In this eye opening book he shows how a solar revolution is developing based on one of Einstein's lesser known discoveries one that gave us laptop computers and mobile phones An accessible guide to renewable technology and a hard hitting critique of the arguments of solar sceptics The Burning Answer outlines a future in which the fuel for electric cars will be generated on our rooftops It is above all an impassioned call to arms to join the solar revolution before it's too late      **Modern Inorganic Chemistry** J. J. Lagowski, 1973

*Supplementary Readings for Chemical Bond Approach* Robert K. Fitzgerald, William Franklin Kieffer, 1960      British Chemical Abstracts, 1932      *Journal of the American Chemical Society* American Chemical Society, 1964 Proceedings of the Society are included in v 1 59 1879 1937      *British Abstracts*, 1953      **American Druggists' Circular and Chemical Gazette**, 1878      **The Lancet**, 1887      **Journal of the Society of Chemical Industry** Society of Chemical Industry (Great Britain), 1890

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