

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|------------------------------------|
| acid | acid dissociation constant (K_a) | acid ionization constant (K_a) |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
| $pH = -\log[H^+]$ (Equation 19.3) | $pOH = -\log[OH^-]$ (Equation 19.4) |

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing the correct number of atoms of each element in the formula.

1. Hydrochloric acid, HCl
2. Acetic acid, CH_3COOH
3. Nitric acid, HNO_3
4. Sulfuric acid, H_2SO_4
5. Phosphoric acid, H_3PO_4
6. Hydrofluoric acid, HF
7. Hydrobromic acid, HBr
8. Hydroiodic acid, HI
9. Hydrocyanic acid, HCN
10. Hydroperfluoroboric acid, $HFBO_2$
11. Hydroperfluoroboric acid, $HFBO_2$
12. Hydroperfluoroboric acid, $HFBO_2$
13. Hydroperfluoroboric acid, $HFBO_2$
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**Frederic Martini, William C. Ober, Edwin
F. Bartholomew**

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