

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. K_a | 9. K_b |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include proper units when appropriate.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of those species present in water that equal to _____.
3. _____ is the concentration of those species present in water that equal to _____.
4. _____ is the concentration of those species present in water that equal to _____.
5. _____ is the concentration of those species present in water that equal to _____.
6. _____ is the concentration of those species present in water that equal to _____.
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15. _____ is the concentration of those species present in water that equal to _____.
16. _____ is the concentration of those species present in water that equal to _____.
17. _____ is the concentration of those species present in water that equal to _____.
18. _____ is the concentration of those species present in water that equal to _____.
19. _____ is the concentration of those species present in water that equal to _____.
20. _____ is the concentration of those species present in water that equal to _____.

1. _____
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20. _____

19 2 Hydrogen Ions And Acidity Answer Key

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