

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on  $K_a$  and  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                  |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain                | 6. acid rain     |
|                   | 7. acid solution            | 8. pH            |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the quantity indicated that are indicated in the problem. Be sure to include units where appropriate.

1. What substance is \_\_\_\_\_, a strong hydrogen ion donor?
2. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions present in 100% pure water at \_\_\_\_\_.
3. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in \_\_\_\_\_.
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# 19 2 Hydrogen Ions And Acidity Section Review Answers

**Kevin T. Patton, Frank B. Bell, Terry  
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