

Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



- a) If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reagent
- b) determine the number of moles of carbon dioxide produced
- c) determine the number of grams of H_2O produced
- d) determine the number of grams of excess reagent left

2. Given the following equation:



- a) If 10.0 g of $\text{Al}_2(\text{SO}_4)_3$ is reacted with 10.0 g of NaOH , determine the limiting reagent
- b) Determine the number of moles of $\text{Al}(\text{OH})_3$ produced
- c) Determine the number of grams of Na_2SO_4 produced
- d) Determine the number of grams of excess reagent left over in the reaction

3. Given the following equation:



- a) If 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe , determine the limiting reagent
- b) Determine the number of moles of Al produced
- c) Determine the number of grams of Fe_3O_4 produced
- d) Determine the number of grams of excess reagent left over in the reaction

Answer Key To Limiting Reagent

Darrell D. Ebbing

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